

Chemistry Benchmark II Study Guide

1. Write the formula for the following:
  - iron (II) oxide
  - calcium phosphate
  - ammonium phosphate
  - copper (III) bromide
  - nitrogen dioxide
  - titanium (IV) oxide
  - potassium bromide
  - magnesium hydroxide
  - xenon hexafluoride
2. What is a binary compound? What is a ternary compound?
3. Write formulas for:
  - hydroiodic acid
  - carbonic acid
  - hydrosulfuric acid
  - hydrochloric acid
  - hypochlorous acid
  - perchloric acid
  - phosphorous acid
  - chloric acid
  - sulfuric acid
  - chlorous acid
4. Name these:
  - NaCl
  - KI
  - CaBr
  - Rb<sub>2</sub>O
  - Sr<sub>2</sub>
  - CuCl
  - HgO
  - Fe<sub>2</sub>O<sub>3</sub>
  - MnO<sub>2</sub>
  - PbCl<sub>4</sub>
  - SnBr<sub>4</sub>
  - BF<sub>3</sub>
  - NO
  - N<sub>2</sub>O<sub>5</sub>
  - CCl<sub>4</sub>
  - P<sub>2</sub>O<sub>5</sub>
  - N<sub>2</sub>O<sub>3</sub>
5. Find the oxidation number assigned to:
  - S in H<sub>2</sub>SO<sub>4</sub>
  - P in PCl<sub>5</sub>
  - S in SO<sub>3</sub>
6. Name these:
  - NO<sub>2</sub>
  - N<sub>2</sub>O
  - SO<sub>3</sub>
7. How many valence electrons do elements have in the following groups?
  - alkali metals
  - alkaline earth metals
  - halogens
  - noble gases
8. Write the complete electron configuration, noble gas configuration, Lewis dot structure, and orbital notation for:
  - Sr
  - Br
  - Na
  - K
  - N
  - Cl
9. In each group, identify the atom with the largest atomic radius.
  - Sn, Xe, Rb, Sr
  - IRn, He, Xe, Kr
  - Pb, Ba, Cs, At
10. In each group, identify the atom with the highest ionization energy.
  - Cs, K, Li
  - Ba, Sr, Ca
  - I, Br, Cl
  - Mg, Si, S
11. Define the following: ionic bond, covalent bond, polar covalent bond, nonpolar covalent bond.
12. Define electronegativity and describe the trend across a period and down a group. What type of element has the highest electronegativity? What type has the lowest electronegativity?
13. Identify the type of bond in each:
  - K-N
  - Cs-O
  - C-N
  - O-F
  - Na-Cl
  - Cl-F
  - Cl-O
  - Ca-Cl
14. Who first arranged the periodic table in order of increasing atomic mass?
15. Draw Lewis structures for the following and predict the molecular shape.
  - H<sub>2</sub>S
  - PH<sub>3</sub>
  - O<sub>2</sub>
  - sulfate ion
  - SO<sub>3</sub>
  - BF<sub>3</sub>
16. Calculate the wavelength and energy of 96.6MHz radio waves.
17. What is the frequency and energy of electromagnetic radiation having a wavelength of  $3.33 \times 10^{-9}$ m?
18. Predict the ions that the following elements might form:
  - F
  - K
  - I
  - Ge
  - Fe (trick question)
  - P
  - Ag

# Chemistry Benchmark Study Guide

**Nader Rifai**

## **Chemistry Benchmark Study Guide:**

Home and School Reading and Study Guides Donna M. Lusardi, Rosemarie Kent, 2007      **Subject Guide to Books in Print**, 2001      **Technical Abstract Bulletin**, 1978      **Tietz Textbook of Laboratory Medicine - E-Book** Nader Rifai, 2022-02-03 Use THE definitive reference for laboratory medicine and clinical pathology Tietz Textbook of Laboratory Medicine 7th Edition provides the guidance necessary to select, perform and evaluate the results of new and established laboratory tests. Comprehensive coverage includes the latest advances in topics such as clinical chemistry, genetic, metabolic disorders, molecular diagnostics, hematology and coagulation, clinical microbiology, transfusion medicine and clinical immunology. From a team of expert contributors led by Nader Rifai, this reference includes access to wide ranging online resources on Expert Consult featuring the comprehensive product with fully searchable text, regular content updates, animations, podcasts over 1300 clinical case studies, lecture series and more. Authoritative current content helps you perform tests in a cost effective, timely and efficient manner, provides expertise in managing clinical laboratory needs and shows how to be responsive to an ever changing environment. Current guidelines help you select, perform and evaluate the results of new and established laboratory tests. Expert internationally recognized chapter authors present guidelines representing different practices and points of view. Analytical criteria focus on the medical usefulness of laboratory procedures. Use of standard and international units of measure makes this text appropriate for any user anywhere in the world. Elsevier eBooks provides the entire text as a fully searchable eBook and includes animations, podcasts, more than 1300 clinical case studies, over 2500 multiple choice questions, a lecture series and more, all included with print purchase. NEW 19 additional chapters highlight various specialties throughout laboratory medicine. NEW Updated peer reviewed content provides the most current information possible. NEW The largest ever compilation of clinical cases in laboratory medicine is included with print purchase on Elsevier eBooks. NEW Over 100 adaptive learning courses included with print purchase on Elsevier eBooks offer the opportunity for personalized education.

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