

In your textbook, read about molarity and preparing molar solutions.

Read the following problem and then answer the questions.

An 85.0-mL aqueous solution contains 7.54 g iron(II) chloride (FeCl_2). Calculate the molarity of the solution.

5. What is the mass of the solute?

7.54 g

6. What is the volume of the solution?

85.0 mL solution

7. Write the equation that is used to calculate molarity.

$$M = \frac{\text{moles of solute}}{\text{L solution}}$$

8. In what unit must the amount of the solute be expressed to calculate molarity?

moles

9. In what unit must the volume of the solution be expressed to calculate molarity?

Liter

10. Write the expression needed to convert the volume of the solution given in the problem to the volume needed to calculate molarity.

$$85.0 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 0.0850 \text{ L}$$

Lesson 13.2 continued

11. What quantity must be used to convert the mass of the solute given in the problem to the amount calculate molarity?

molar mass of FeCl_2 126.75 g/mol

12. Write the expression used to calculate the amount of solute

$$7.54 \text{ g FeCl}_2 \times \frac{1 \text{ mol}}{126.75 \text{ g}} = 0.0595 \text{ mol}$$

13. Calculate the molarity of the solution. Show all your work.

$$\underbrace{7.54 \text{ g} \times \frac{1 \text{ mol}}{126.75}}_{\text{moles of Solute}} \times \underbrace{\frac{1}{0.0850 \text{ L}}}_{\text{L of Solution}} = \textcircled{0.700 \text{ M}}$$

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Chemistry Chapter 13 Study Guide

Ralph J. Fessenden, Joan S. Fessenden



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Table of Contents Chemistry Chapter 13 Study Guide

1. Understanding the eBook Chemistry Chapter 13 Study Guide
 - The Rise of Digital Reading Chemistry Chapter 13 Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 13 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 13 Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 13 Study Guide
 - Personalized Recommendations
 - Chemistry Chapter 13 Study Guide User Reviews and Ratings

- Chemistry Chapter 13 Study Guide and Bestseller Lists
- 5. Accessing Chemistry Chapter 13 Study Guide Free and Paid eBooks
 - Chemistry Chapter 13 Study Guide Public Domain eBooks
 - Chemistry Chapter 13 Study Guide eBook Subscription Services
 - Chemistry Chapter 13 Study Guide Budget-Friendly Options
- 6. Navigating Chemistry Chapter 13 Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 13 Study Guide Compatibility with Devices
 - Chemistry Chapter 13 Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 13 Study Guide
 - Highlighting and Note-Taking Chemistry Chapter 13 Study Guide
 - Interactive Elements Chemistry Chapter 13 Study Guide
- 8. Staying Engaged with Chemistry Chapter 13 Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 13 Study Guide
- 9. Balancing eBooks and Physical Books Chemistry Chapter 13 Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 13 Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 13 Study Guide
 - Setting Reading Goals Chemistry Chapter 13 Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 13 Study Guide
 - Fact-Checking eBook Content of Chemistry Chapter 13 Study Guide
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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