

CHAPTER 17 **STUDY GUIDE****Chemical Equilibrium****Section 17.1 Equilibrium: A State of Dynamic Balance***In your textbook, read about chemical equilibrium.***Complete each statement.**

1. When a reaction results in almost complete conversion of reactants to products, chemists say the reaction goes to _____.
2. A reaction that can occur in both the forward and the reverse directions is called a(n) _____.
3. _____ is a state in which the forward and reverse reactions balance each other because they take place at equal rates.
4. At equilibrium, the concentrations of reactants and products are _____ but that does not mean that the amounts or concentrations are _____.
5. Equilibrium is a state of _____, not one of _____.

*In your textbook, read about equilibrium expressions and constants.***For each statement below, write true or false.**

6. The law of chemical equilibrium states that at a given pressure, a chemical system may reach a state in which a particular ratio of reactant to product concentrations has a constant value.
7. The equation $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$ is an example of a homogeneous equilibrium.
8. If an equilibrium constant has a value less than one, the reactants are favored at equilibrium.
9. The value for K_{eq} is constant only at a specific volume.
10. If the equilibrium constant for a reaction at 300 K is 49.7, the concentration of the reactants will be greater than the concentration of the products.
11. A heterogeneous equilibrium means that reactants and products are present in more than one state.
12. The product of the forward chemical reaction is HI , for the equilibrium expression:

$$K_{\text{eq}} = \frac{[\text{HI}]^2}{[\text{H}_2][\text{I}_2]}$$

Chemistry Chemical Equilibrium Study Guide Key

Donald J. DeCoste

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Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2002-01-04 This third edition continues to innovate by providing students with an integrated and modern approach to the subject. The text emphasizes the modern tools of chemistry while incorporating historical evidence and its unique molecular quantitative emphasis is further reinforced by an integrated media package developed by the authors. Also of benefit is the just in time presentation of key content only providing details once they are needed. While key topics and analytical techniques have been updated there is now an additional third chapter on chemical equilibrium. The authors have also developed an expanded and more integrated problem solving emphasis that now incorporates a 4 step strategy throughout complete with text icons. The whole is backed by a range of supplements including a new illustration program a tutorial CD interactive learningware an extensive Web CT component an instructor's resource CD and a solution CD [Study Guide for Chemical Principles \[by\] Steven S. Zumdahl](#) Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text. The new edition of the Study Guide includes many new worked out examples

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