

Study Guide for Chapter 7 – Chemical Quantities

Test Setup: Multiple Choice, Matching, Problems: Find molar mass, determine % composition, Molar conversions will include all of the following: moles-particles, moles-mass, and moles-volume, determine the empirical formulas & molecular formulas by looking at actual formulas.

Sections in the Book:

7.1 – The Mole: A Measurement of Matter

Objectives:

- Describe how Avogadro's number is related to a mole of any substance
- Calculate the mass of a mole of any substance

Vocabulary:

- mole (mol), Avogadro's number, representative particle

Practice Problems:

1. Find the molar mass of Li_2S .
2. Find the molar mass of $\text{Ca}(\text{OH})_2$.
3. How many moles in 1.50×10^{23} molecules NH_3 .
4. How many formula units in 6.78 moles of sodium chloride (NaCl)

7.2 – Mole-Mass and Mole-Volume Relationships

Objectives:

- Use the molar mass to convert between mass and moles of a substance.
- Use the mole to convert among measurements of mass and volume of substances.

Vocabulary:

- molar mass, standard temperature and pressure (STP), molar volume

Practice Problems:

1. What is the mass in grams of 0.160 mol H_2O_2 ?
2. How many moles are in 508.4 g $\text{Ca}(\text{NO}_3)_2$?
3. Calculate the volume, in liters, of 8.75 mol of chlorine gas at STP.
4. How many moles are in a container filled with 45.9 L of oxygen gas at STP?

7.3 – Percent Composition and Chemical Formulas

Objectives:

- Calculate the percent composition of a substance from its chemical formula or experimental data.
- Derive the empirical formula and the molecular formula of a compound from experimental data.

Chemistry Chemical Quantities Study Guide

Richard H. Eastman

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Study Guide to Accompany Basics for Chemistry Martha Mackin,1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students **Study Guide/Selected Solutions Manual** Julie R. Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book **Student Study Guide to Accompany Chemistry** Kenneth J. Hughes,Paul B. Kelter,2002-05 Reviews essential points from the text presents problems and hints to assist the student in solving them Solutions to selected problems in the

end of chapter problem sets are included By Kenneth Hughes of the University of Wisconsin Madison *Solutions Guide for Introductory Chemistry* James F. Hall,1996 **The Publishers' Trade List Annual** ,1978 **Study Guide and Solutions Manual for Fundamentals of General, Organic, and Biological Chemistry** Susan McMurry,1992 *Chemical Abstracts* ,1921 **Solutions Guide, Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry, Fourth Edition, Zumdahl** James F. Hall,2000 *Fundamentals of Chemistry* Burns A. Ralph,Charles Baker,Claire Baker,1995 This book is a supplement to the texts not a replacement It is intended to maximize your success in this course by showing you how to become involved in developing your own techniques for grasping the concepts of chemistry Using a study outline problem sets problem examples worked and unworked and numerous self tests with answers this manual will provide you with opportunities to sharpen your skills and evaluate your comprehension of the material in the texts The worked out solutions at the end of this manual walk you step by step through the methods of arriving at the answer to those same problems which have an answer only in the answer key Used in conjunction with your textbook and classroom lecture notes this book offers an essential learning opportunity to the chemistry student Back cover *Chemistry* Zumdahl,Steven S. Zumdahl,2003-07 Contains chapter discussions key term definitions and practice chapter tests

General Chemistry: Experiment and Theory Richard H. Eastman,1970 **The Chemical Trade Journal and Chemical Engineer** G Kelville Davis,1928 **Platers' Guide** ,1921 *The Chemical Review* ,1884 **Study Guide for Chemical Principles [by] Steven S. Zumdahl** Steven S. Zumdahl,Paul B. Kelter,1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

Embracing the Melody of Term: An Mental Symphony within **Chemistry Chemical Quantities Study Guide**

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