



CHEMISTRY CHEAT SHEET



EQUILIBRIUM

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Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol K_c .

What is Equilibrium?
 A reversible reaction is one that can proceed in either direction. At equilibrium, the rate of the forward reaction is equal to the rate of the reverse reaction. The concentrations of the reactants and products remain constant over time.

Equilibrium Constant K_c
 The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of their stoichiometric coefficient. It is a constant for a given reaction at a given temperature.



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Reaction	Equilibrium Constant K_c
$aA + bB \rightleftharpoons cC + dD$	$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$
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Factor	Effect on Equilibrium	Example
Concentration	Increasing the concentration of a reactant or product shifts the equilibrium in the direction that consumes that substance.	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing $[N_2]$ or $[H_2]$ shifts the equilibrium to the right, producing more NH_3 .
Pressure	Increasing the pressure shifts the equilibrium in the direction that produces fewer moles of gas.	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing pressure shifts the equilibrium to the right, producing more NH_3 .
Temperature	Increasing the temperature shifts the equilibrium in the direction that absorbs heat (endothermic reaction).	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing temperature shifts the equilibrium to the left, producing less NH_3 .

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Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

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