

## Chemistry Final Exam Review Questions

### Useful Values and Equations

Some of the following values and equations may be useful for answering some of the questions and problems.

$$\text{op} = 1.84(\text{°C}) + 32$$

$$\text{speed of light} = 3.00 \times 10^8 \text{ m/s}$$

$$c = \lambda\nu$$

$$R = 0.082057 \text{ L-atm / deg-mole}$$

$$\text{Planck's constant} = 6.63 \times 10^{-34} \text{ J·s}$$

$$E = hv$$

### Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- One of the following does not describe gases. Which one is it?
  - Expands infinitely.
  - Rigid shape.
  - Easily compressed.
  - Relatively low densities.
- The following properties describe the element, mercury (Hg). Which one is a **chemical** property?
  - Its density is 5.43 g/cm<sup>3</sup>.
  - It is a silvery-white metal.
  - It is a liquid at 298 K.
  - Mercury(II) sulfide forms when it is exposed to sulfide ions.
- All of the following are properties of SO<sub>2</sub>. Which one is a **physical** property?
  - It reacts with oxygen to form SO<sub>3</sub><sup>2-</sup>.
  - It condenses to a colorless liquid at 14°F and atmospheric pressure.
  - It is a strong reducing agent.
  - It reacts with water and oxygen to form an acid.
- Which of the following numbers has 4 significant figures?
  - 0.04309
  - 0.0430
  - 0.0431
  - 0.43900
- How many significant figures should there be in the answer to the following math equation?  
 $2.834 + 5.71520 + 2.12 + 178.1 + 250.2619$ 
  - 3
  - 4
  - 5
  - 6
- How many significant figures should there be in the answer to the following math equation?  
 $23.413 \text{ g} + (2.15 \text{ cm} \times 1.1 \text{ cm} \times 3.75 \text{ cm})$ 
  - 2
  - 3
  - 4
  - 5
- At 25°C, one cm<sup>3</sup> of iron has a mass of 7.62 grams. What volume of iron is occupied by 12.5 grams of iron?
  - 0.82 cm<sup>3</sup>
  - 8.2 cm<sup>3</sup>
  - 0.164 cm<sup>3</sup>
  - 16.4 cm<sup>3</sup>
- A cube of lead has a volume of 6.61 cm<sup>3</sup>. Its mass is 75.0 grams. What is its density?
  - 3.85 g/cm<sup>3</sup>
  - 7.41 g/cm<sup>3</sup>
  - 11.3 g/cm<sup>3</sup>
  - 22.6 g/cm<sup>3</sup>

# Chemistry Final Exam Spring 2014 Answers

**A Gutmann**

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