

Chemistry MSL Review

Math in Chemistry

- Give a definition and example of a qualitative measurement. Give a definition and example of a quantitative measurement.
- How many sig figs are contained in the following numbers:
0.067 2450 2450.0 0.0100
- Define accuracy and precision.
- If a block has a known density of 1.45, what is the percent error of the value 1.14 measured in lab?
- A block of unknown composition has a mass of 23.2 grams. Using water displacement, its volume is measured at 27.9 mL. What is the block's density?
- What is the volume of an object with a density of 1.94 g/mL and a mass of 15.5g?
- Convert the temperature 88°C to Kelvin.
- Metric conversions: Convert 4.56 m to cm. Convert 304.5 J to kJ. Convert 0.009 332 g to mg.
- What is the difference between a chemical change and a physical change?

Atomic Structure

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge; too few, a positive charge.

This structure can be written as part of a chemical symbol.

Example:	mass number + $^{15}_{7}\text{N}^{+3}$ atomic number	charge: 7 protons 8 neutrons ($15 - 7$) 4 electrons
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- Define the following terms:
atomic number; mass number; isotope; atomic mass
- How do you determine the number of protons in an atom? How do you determine the number of neutrons in an atom?
- Complete the following table:

Isotope	Atomic #	# Protons	# Neutrons	Mass #
^{12}C				
	17		18	
		26		56
		17		37
			2	3
	52			126
		50	70	

- Consider the second and fourth row in the table above. What do they have in common?
- Consider the second and fourth row in the table above. What are the differences?

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