

Chemical Monitoring and Management

9.4.1 Much of the work of chemists involves monitoring the reactants and products of reactions and managing reaction conditions

Outline the role of a chemist employed in a named industry or enterprise, identifying the branch of chemistry undertaken by the chemist and explaining a chemical principle that the chemist uses
Gather, process and present information from secondary sources about the work of practising scientists identifying:

- the variety of chemical occupations
- a specific chemical occupation for a more detailed study

An analytical chemist working for Sydney Water will use AAS to monitor concentrations of metals, such as Pb and Hg, in water samples which eventually will be supplied to households. The principle of chemistry used by the scientist is that metal electrons move from one energy shell to a higher one by absorbing electromagnetic radiation of a wavelength specific to that metal, giving each metal a unique emission spectrum $\therefore$ metals can be identified and their concentration can be determined even in the presence of other metals. It is important that the scientist monitor and regulate the concentrations of metals in drinking water sources as excessive concentrations can lead to poisoning and illness, e.g. lead poisoning can cause mental retardation, but trace element metals, like Zn, are essential in small concentrations for body functioning.

Identify the need for collaboration between chemists as they collect and analyse data

- Chemists specialising in various fields of chemistry, e.g. analytical, organic, industrial, need to collaborate to solve problems requiring a broad spectrum of chemical knowledge.
- Chemists need to regularly exchange viewpoints and have an open minded but critical approach to ensure that our scientific knowledge is constantly improving.

Describe an example of a chemical reaction such as combustion, where reactants form different products under different conditions and thus would need monitoring

Combustion reactions can produce solely carbon dioxide or a mixture of poisonous carbon monoxide and pollutant soot (C), depending on the amount of oxygen present. Using methane as an example:

Oxygen Supply	Bunsen Burner Hole	Flame Colour	Rate of Combustion	Reaction Equation
Sufficient	Open	Blue	Complete combustion; maximum energy released	$\text{CH}_4 (\text{g}) + 2\text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + 2\text{H}_2\text{O} (\text{l})$
Limited	Partially open	Mauve	Incomplete combustion	$2\text{CH}_4 (\text{g}) + 3\text{O}_2 (\text{g}) \rightarrow 2\text{CO} (\text{g}) + 4\text{H}_2\text{O} (\text{l})$
Blocked	Closed	Yellow	Incomplete combustion	$\text{CH}_4 (\text{g}) + \text{O}_2 (\text{g}) \rightarrow \text{C} (\text{s}) + 2\text{H}_2\text{O} (\text{l})$

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This volume is an important study for understanding the complex interconnections between basic science and its sources of economic support in the period between the two world wars. The focus of the study is on the Institute for Theoretical Physics later renamed the Niels Bohr Institute at Copenhagen University and the role of its director the eminent Danish physicist Niels Bohr in the funding and administration of the Institute. Under Bohr's direction the Copenhagen Institute was a central workplace in the development and the formulation of quantum mechanics in the 1920s and later became an important center for nuclear research in the 1930s. Dr Aaserud brings together the scholarship on the internal origins and development of nuclear physics in the 1930s with descriptions of the concurrent changes in private support for international basic science particularly as represented by Rockefeller Foundation philanthropy. In the process the book places the emergence of nuclear physics in a larger historical context. This book will appeal to historians of science, physicists and advanced students in these areas. **Excel Chemistry** C. M. Roebuck, 1995 *A+ HSC Year 12 Chemistry Study Notes* Colin Harrison, 2022 *Excel HSC Chemistry* Jim Stamell, 2011 ISBN 9781741252996 AUTHOR Jim Stamell RRP 39.95 PAGES 428 pp SPECIFICATION

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