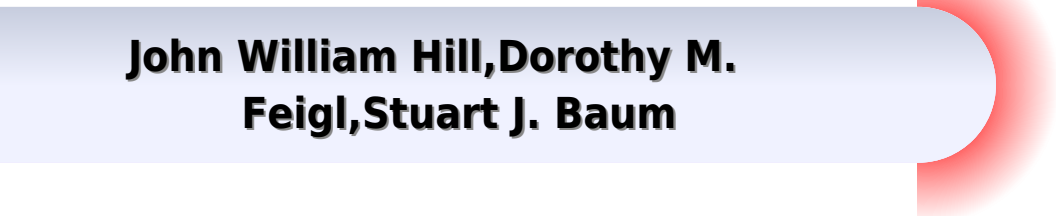


Chapter 2 – The Chemistry of Life Study Guide

1. A solution with a high concentration of H ions is (very acidic/very basic).
2. Chemical reactions that absorb more energy than they release are (endergonic/exergonic).
3. The first carbon compounds that scientists studied are called _____ compounds because they came from living compounds.
4. DNA and RNA are two types of _____ (Look in your notes at the 4 Major Classes of Organic Molecules).
5. What is the process that changes one set of chemicals into another set of chemicals?
6. When salt is dissolved in water, water is the (solvent/solute)?
7. Enzymes affect chemical reactions in living organisms by
 - a. Changing the direction of a reaction
 - b. Breaking down molecules into starch
 - c. Weakening bonds in reactants
 - d. Increasing the temperature range
8. How many atoms of oxygen and hydrogen make up a molecule of water?
9. _____ cannot be broken down by ordinary chemical means.
10. _____ bonds hold molecules together.
11. The three particles that make up an atom are _____.
12. What type of ion forms when an atom loses electrons (positive/negative)?
13. _____ represents a formula for a chemical compound.
 - a. H
 - b. H₂O
 - c. C
 - d. P
14. Which of the following solutions has the highest H ion concentration?
 - a. A solution with a pH of 7
 - b. A solution with a pH of 1
 - c. A solution with a pH of 4
 - d. A solution with a pH of 10
15. What are the four main groups of carbon-based molecules?
16. The activation energy needed for a chemical reaction is decreased by a _____.
17. _____ describes a substance formed by the combination of two or more elements in definite proportions.

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**John William Hill, Dorothy M.
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