

# AP Chemistry Pacing Guide

<ul style="list-style-type: none"> <li>Approximation</li> <li>Multistep Reaction</li> <li>Energy Profile</li> <li>Catalysis</li> </ul>		<p>first step is not rate limiting.</p> <p>10. I can represent the activation energy and overall energy change in a multistep reaction with a reaction energy profile.</p> <p>11. I can explain the relationship between the effect of a catalyst on a reaction and changes in the reaction mechanism.</p>	
<b>Unit 6: Thermodynamics</b> <ul style="list-style-type: none"> <li>Endothermic and Exothermic Processes</li> <li>Energy Diagrams</li> <li>Heat Transfer and Thermal Equilibrium</li> <li>Heat Capacity and Calorimetry</li> <li>Energy of Phase Changes</li> <li>Introduction to Enthalpy of Reaction</li> <li>Bond Enthalpies</li> <li>Enthalpy of Formation</li> <li>Hess's Law</li> </ul>	<b>BIG IDEA 4:</b> Energy	<ol style="list-style-type: none"> <li>I can explain the relationship between experimental observations and energy changes associated with a chemical or physical transformation.</li> <li>I can represent a chemical or physical transformation with an energy diagram.</li> <li>I can explain the relationship between the transfer of thermal energy and molecular collisions.</li> <li>I can calculate the heat <math>q</math> absorbed or released by a system undergoing heating/cooling based on the amount of the substance, the heat capacity, and the change in temperature.</li> <li>I can explain changes in the heat absorbed or released by a system undergoing a phase transition based on the amount of the substance in moles and the molar enthalpy of the phase transition.</li> <li>I can calculate the heat <math>q</math> absorbed or released by a system undergoing a chemical reaction in relation to the amount of the reacting substance in moles and the molar enthalpy of reaction.</li> <li>I can calculate the enthalpy change of a reaction based on the average bond energies of bonds broken and formed in the reaction.</li> <li>I can calculate the enthalpy change for a chemical or physical process based on the standard enthalpies of formation.</li> <li>I can explain the relationship between the enthalpy of a chemical or physical process and the sum of the enthalpies of the individual steps.</li> </ol>	10 days
<b>Unit 7: Equilibrium</b> <ul style="list-style-type: none"> <li>Introduction to Equilibrium</li> <li>Direction of Reversible Reactions</li> <li>Reaction Quotient and Equilibrium Constant</li> <li>Calculating the Equilibrium Constant</li> <li>Magnitude of the Equilibrium Constant</li> </ul>	<b>BIG IDEA 1:</b> Scale, Proportion, and Quantity <b>BIG IDEA 3:</b> Transformation	<ol style="list-style-type: none"> <li>I can explain the relationship between the direction in which a reversible reaction proceeds and the relative rates of the forward and reverse reactions.</li> <li>I can represent the reaction quotient <math>Q_c</math> or <math>Q_p</math> for a reversible reaction, and the corresponding equilibrium expressions <math>K_c = Q_c</math> or <math>K_p = Q_p</math>.</li> <li>I can calculate <math>K_c</math> or <math>K_p</math> based on experimental observations of concentrations or pressures at equilibrium.</li> <li>I can explain the relationship between very large or very small values of <math>K</math> and the relative concentrations of chemical species at equilibrium.</li> <li>I can represent a multistep process with an overall equilibrium expression, using the constituent <math>K</math> expressions for each individual</li> </ol>	9 days

# Chemistry Pacing Guide

**Brig Leane, Jon Yost**



## Chemistry Pacing Guide:

The 4 X 4 Block Schedule J. Allen Queen, Kimberly Gaskey Isenhour, 1998 This practical book shows you how to motivate and train teachers establish community support develop new classroom strategies and measure success of the 4 X 4 block schedule

**Educational Curricula** Jennifer N. Casey, 2008 In formal education a curriculum plural curricula is the set of courses and their content offered at a school or university As an idea curriculum stems from the Latin word for race course referring to the course of deeds and experiences through which children grow and mature in becoming adults Crucial to the curriculum is the definition of the course objectives that usually are expressed as learning outcomes and normally include the program s assessment strategy These outcomes and assessments are grouped as units or modules and therefore the curriculum comprises a collection of such units each in turn comprising a specialised specific part of the curriculum So a typical curriculum includes communications numeracy information technology and social skills units with specific specialised teaching of each This book presents research on educational curricula from around the world

**Holt Chemistry** Ralph Thomas Myers, 2004

**Singletons in a PLC at Work®** Brig Leane, Jon Yost, 2022-06-21 In a professional learning community isolation is the enemy of school improvement But what does collaboration among teachers look like when you can t easily identify with a team This book will help singleton teachers first develop clarity on learning essentials then find creative entry points to form collaborative teams Drawing from their own experiences the authors offer practical solutions for eliminating the practice of isolation for all educators Collaborative teams will Understand what meaningful collaboration is and how singletons can utilize the PLC process Build the groundwork for meaningful collaboration using strategies for your specific situation Implement meaningful collaboration as a singleton across separate schools or within the same school Align disparate singletons under the same unifying PLC process Contents Chapter 1 Meaningful Collaboration Chapter 2 Singleton On Ramps for Collaboration Chapter 3 Preparation for Meaningful Collaboration Chapter 4 Course Alike Entry Point The Virtual Team Chapter 5 Common Content Entry Point Chapter 6 Critical Friend Entry Point Chapter 7 Putting It All Together Afterword Final Thoughts References and Resources Index

PSI Study Guide for Organic Chemistry by Douglas C. Neckers, Michael P. Doyle E. C. Blosser, 1977

**Ooey Goey Chemistry** Marie Ecker, 2017-08-15 hands on and active 9 week curriculum designed for homeschool families and Co op classes Student directed activities and learning Labs that can be done at home without any special equipment using standard measuring tools Video alternatives for all labs and demos Pacing guide and teacher tips for all lessons

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