

# Chemistry Study Guide

Periodic Table of the Elements																		
H	Li	Be																
Na	Mg																	
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	
Cs	Ba	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn		
Fr	Ra	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo		
		La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu		
		Ac	Th	Pa	U	Ng	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr		

## Objectives:

To understand the structure of an atom, and the nature + properties of subatomic particles.



## Matter:

Anything that has mass and occupies space.

Examples: air, gases, liquids, solidz!

## Atoms:

Atoms are the smallest, indivisible particle.

- Outer part → shell/orbit/energy level → carries the electronz (e-)
- Inner part → nucleus → houses the protonz (p+) and neutronz (n)
- Outer & inner part → together these are the subatomic particles.

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p <sup>+</sup>	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e <sup>-</sup>	-0	-1	Orbit/Shell

# Chemistry Slo Study Guide

**Karen C. Timberlake**

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*Introductory Chemistry* Mike Lucarelli,David Alan Proctor,2002 *Peterson's Annual Guides to Graduate Study* ,1977

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