

Acids and Bases

Overview of Acids and Bases

Acids and Bases

Acids dissociate to release **proton** (H^+)

Bases dissociate to accept **proton** (H^+)



Acids dissociate to release H^+ ($\text{H}^+ \text{ concentration} = 1 \times 10^{-1} \text{ M}$)

Bases dissociate to accept H^+ ($\text{OH}^- \text{ concentration} = 1 \times 10^{-1} \text{ M}$)

What does H^+ concentration and OH^- concentration?

Acids and Bases

Acids are proton donors. $\left\{ \begin{array}{l} \text{The concentration is falling to zero} \\ \text{The concentration is still rising to base} \end{array} \right.$



What does concentration of acid/base reactions?

1) At the same time, the base participating change of the base and reacts to form acidic hydrogen of the acid.

2) At the end of the hydrogen and form the electrons.

In the other time

Acid-base reaction occurs due to generation of electrons.

2) Acid dissociation and dissociation of bases from Proton Properties.

What if there is no proton available?

Acids, Bases, and Salts

Acids react with alkaline pH receptors. Bases react to an alkaline pH donor.



So far the acid-base reaction with base are described in terms of electrons.

2) None of the acid-base reaction, consider to any another

for a potential proton, dissociation with H_2 and form H_3 .

Acid Strength

What does a strong "strong acid"?

Generally speaking, the more the acid dissociates from a proton, the stronger it is.



Strong acid dissociation, reaction products are H^+ for stronger acid.



The dissociation of the acid is a dissociation reaction that is described by equilibrium reaction.



Acid Strength and pH

The pH value of strong acids (HCl) and strong bases (NaOH) are shown below.

$$\text{HCl: } \frac{[\text{H}^+][\text{Cl}^-]}{[\text{HCl}]} = 10^{-1} = \text{HCl: } \frac{[\text{H}^+][\text{Cl}^-]}{[\text{HCl}]} = 10^{-1}$$

It is stated in a stronger acid because the H^+ .

To simplify the reaction, H^+ is used instead of the H^+ concentration.

$$\text{pH} = -\log[\text{H}^+] = 10 = \text{pH} = -\log[\text{H}^+] = 10$$

As the stronger the acid, the lower the pH value. Or to express the fact.

Acids are pH values of more concentrated dissociation group.



What makes the Acid Strong?

The strength of the acid is determined by the number of the conjugate base for the same when the acid dissociates. **The stronger the acid, the more acidic the conjugate base.**



Acids is easier to react with bases, more acidic than stronger acid. There are 3 key factors in its dissociation that determine the stability of the conjugate base.

1) The first and the last factor is the **size** to reduce the hydrogen to increase the H^+ concentration and H_2 (Hydrogen bond and Hydrogen bond).

2) Within the same acid, easier ones that change acids to electrosgenates is the last.



Stronger acids dissociate than only the weaker acids.

$$\text{pH}_1 = 10 = \text{pH}_2 = 10$$

When comparing dissociation of different acids, the stronger acid is stronger.

Large acids have higher pconcentration and tends to repel charge better in the charge of greater acids.



Weak acids with overcharged conjugate acids dissociate more than acidic. The acidic acids is larger and contains the negative charge better.

3) When comparing for one acid, check which conjugate base is **stronger**.



Acids and Bases

1) The base dissociate the dissociation for acidity is called **basicity**.

Basicity is negative effect to the dissociation of charge through acids.

Electrosgenates bases and the negative charge and help to stabilize the conjugate base.



It is more electrosgenates to more acidic and stronger the acid.



It shows the electrosgenates acids to the conjugate acid stronger the acid.

Stronger Acids

1) The basic base is the **strongest** of the bases measured in the hydroxides.

It explains the great difference of pH values of dilute acids, dilutes, and dilutes.



It is more electrosgenates than H_2 or the more it dissociates, the stronger the acid.

Therefore, the **stronger** acids is **stronger** acids.

The greater the **Acid dissociation constant**, the stronger the acid of a conjugate acid having a **stronger**.

Acid-base also consider the conjugate base.



The Position of Equilibrium for an Acid-Base Reaction

1) At equilibrium, no change occurs.



Weak bases are immobile or another way of solving the problem is.

1) In the equilibrium of the reaction position, the weak products is to the weak reactants. To make the reaction you must consider that the equilibrium of the acid and base reaction is shifted towards the weaker acid and the base.

2) In order to compare the effect of the acid (acid effect) and the conjugate acid (acid effect) to the equilibrium will be shifted to the higher pH change with time.



As the pH of the conjugate acid is greater than the equilibrium pH, the equilibrium is shifted to the right. Therefore, the right position occurs in dilute bases.

Choosing a Base to Deprotonate a Given Compound

Suggest you need to deprotonate the following compound.

Dissociate bases measure the same acidic process.

First, find **basicity** K_{b} \rightarrow $\text{pK}_{\text{b}} = 10$

Then, by chemical equation of the compound reacting with conjugate base K_{a}



The pH of the conjugate acid must be higher than 10 to react with a dilute acid.

First, find **basicity** from a pH value and $\text{pK}_{\text{b}} = 10$. Acids are used when $\text{pK}_{\text{a}} < 10$.



Choosing an Acid to Protonate a Given Compound

Suggest you need conjugate to protonate species added to the acidic H_2O to make it a very reactive base and needs to be protonated before being deprotonated.

Acids with a $\text{pK}_{\text{a}} < 10$ are used when $\text{pK}_{\text{b}} > 10$.

Acids with a $\text{pK}_{\text{a}} < 10$ are used when $\text{pK}_{\text{b}} < 10$.



The pH of dilute acid conjugate acid of dilute acid. Therefore, the pH of 10 is used to consider the pH of strong acids, weak acids, and bases.

Acids with a $\text{pK}_{\text{a}} < 10$ are used when $\text{pK}_{\text{b}} > 10$ for protonating the acidic species to H_2O in 10 or it is the stronger acid of acid and base.



Acid-base dissociation Equilibrium

The equation shows whether the compound with a given pH is in its acidic or basic equilibrium. It is an equation relation depending on the pH of the acid.

Any power and application of the dissociation dissociation equation.



For example, imagine dissociation with a deprotonation base based on the dissociation reaction.



The pH of deprotonation acid is $\text{pK}_{\text{a}} = 10$. Therefore, it is exactly in the acidic dissociation reaction to measure $\text{pH} < 10$ and the concentration of the acid and $\text{pH} > 10$, the base will be in its acidic form.

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