

Acids and Bases

Overview of Acids and Bases

Indicators: Acids and Bases

Acids: Donates its proton to base (proton donor)

Base: Donates its lone pair to acid (LBP)



Remember: (1) only H atoms, (2) only OH is a base

What about organic indicators and indicators solution?

Acid-Base Indicators

Acids are proton donors, Bases are proton acceptors.



General acid-base reactions of acid-base reactions:

to find the strong base, the base gets protonated, change of the base and usually for weak acids hydrogens of the acid.

to find the level of the hydrogens and more the electrons in the other atom.

Remember: -OH is a weak base, -OH is a weak acid.

1. Acid: Donates and Proton: Donates, so: Proton Donor: Protonator. What if there is no proton transfer?

Acid-Base Indicators

Acids: Donates its proton to base (proton donor). Bases: Donates its lone pair to acid (LBP).



Acid-Base Indicators: Acids and Bases are described in terms of electron.

1. Acids: Donates its proton to base (proton donor). Bases: Donates its lone pair to acid (LBP).

Acid Strength

What does it mean "strong acid"?

Indicating: Acids: Donates its proton to base (proton donor). Bases: Donates its lone pair to acid (LBP).



The dissociation of the acid is a reversible reaction that is described by equilibrium constant.



Acid Strength and pKa

The Ka value of an acid (acid) and its conjugate base (B) are shown below.



The acid is a stronger acid than its conjugate base.

To compare the acid, the pKa is used, which is the pKa of the given acid.

$$\text{pKa} = -\log K_a$$

$$\text{pKa}_1 = -\log K_{a1} = 10, \quad \text{pKa}_2 = -\log K_{a2} = 10$$

Acid-Base Indicators: Acids and Bases are described in terms of electron.



What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base. The more stable the conjugate base, the stronger the acid.

The stronger the acid, the more stable the conjugate base.



Acid-Base Indicators: Acids and Bases are described in terms of electron.

There are two factors to be considered when assessing the stability of the conjugate base.

1. The first and the last factor is the **charge** which the hydrogens is concerned. Usually, an **anion** (like **Phenoxide** and **Phenolate**) and **Phenolate** are **stronger** acids.

2. When the acid is **strong**, it is **stronger** than its conjugate base.



Stronger is more electronegative, that is why the acid is more acidic.

$$\text{pKa}_1 = 10, \quad \text{pKa}_2 = 10$$

1. When comparing, always in different ways, the strong acid is better.

Large atoms have higher polarizability and usually for negative charge better in the charge is spread out.



There are acid-base reactions that are more electronegative than acids.

The acid is more stable and usually for negative charge better.

2. When comparing for acid strength, check which conjugate base is **stronger** (stronger).



1. The first factor for determining the acidity is called **inductive**.

Inductive effect is the transmission of charge through atoms.

Electronegativity atoms pull the negative charge and help to stabilize the conjugate base.



The more electronegative, the more electron withdrawing, the stronger the acid.



The closer the electronegative atom to the hydrogens, the stronger the acid.



2. The second factor is the **charge** of the atom connected to the hydrogens.

The electron pull difference of pKa values of acids, alcohols, and aldehydes.



3. Inductive effect: Acids, Bases, and Indicators, Inductive effect: Acids, Bases, and Indicators.

Inductive effect: Acids, Bases, and Indicators, Inductive effect: Acids, Bases, and Indicators.

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Acid-Base Indicators: Acids and Bases are described in terms of electron.

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The Position of Equilibrium for an Acid-Base Reaction

Is the following reaction reaction possible?



Acid-Base Indicators: Acids and Bases are described in terms of electron.

Is the equilibrium of the reaction shifted to the right (products) or to the left (reactants)?

To answer this question, you need to remember: for the equilibrium of acids and bases reaction is shifted towards the weaker acid and base.

For you need to compare the pKa of the acid and base and the conjugate acid and base.

The equilibrium will be shifted to the higher pKa reaction with only.



Is the acid the pKa of the conjugate acid is greater than the acid pKa, or the equilibrium is shifted to the right. Therefore the right reaction is more likely.

Choosing a Base to Deprotonate a Given Compound

Suppose you need to deprotonate the following compound: **CH3COOH**.

Deprotonate means remove the proton (H+).

First, find the pKa value: **CH3COOH** pKa = 4.75.

Find the chemical equation of the conjugate base with a conjugate base (B-).



The pKa of the conjugate acid must be higher than 4.75, so the reaction will be favored.

Find a conjugate base with a pKa value with pKa > 4.75. Acids are weak acids (pKa > 4.75).



Choosing an Acid to Protonate a Given Compound

Suppose you need to protonate the following compound: **CH3COO-**. What is the pKa value of the acid?

Is the acid strong enough to protonate the conjugate base?

To do that, write the chemical equation of the acid with a conjugate base.



The pKa of the acid must be higher than 4.75, so the reaction will be favored.

Find a conjugate base with a pKa value with pKa > 4.75. Acids are weak acids (pKa > 4.75).

Acids are weak acids (pKa > 4.75). Acids are weak acids (pKa > 4.75).



Hammett-Hansch Equation

The equation shows whether the compound with a given pKa is in the acid or base (deprotonated) form in an aqueous solution depending on the pH of the solution.

$$\text{pKa} = \text{pH} + \log \frac{[\text{A}^-]}{[\text{HA}]}$$

For given and application of the Hammett-Hansch equation.

1. If the pH > pKa, 2. If the pH < pKa, 3. If the pH = pKa.

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