

Chemistry Study Guide

Periodic Table of the Elements

<div>1 H</div>	<div>2 He</div>																	<div>18 Ar</div>	<div>19 K</div>	<div>20 Ca</div>											<div>36 Kr</div>	<div>37 Rb</div>	<div>38 Sr</div>											<div>54 Xe</div>	<div>55 Cs</div>	<div>56 Ba</div>															<div>86 Rn</div>																															
<div>3 Li</div>	<div>4 Be</div>																	<div>12 Mg</div>	<div>13 Al</div>	<div>14 Si</div>	<div>15 P</div>	<div>16 S</div>	<div>17 Cl</div>	<div>18 Ar</div>	<div>19 K</div>	<div>20 Ca</div>	<div>21 Sc</div>	<div>22 Ti</div>	<div>23 V</div>	<div>24 Cr</div>	<div>25 Mn</div>	<div>26 Fe</div>	<div>27 Co</div>	<div>28 Ni</div>	<div>29 Cu</div>	<div>30 Zn</div>	<div>31 Ga</div>	<div>32 Ge</div>	<div>33 As</div>	<div>34 Se</div>	<div>35 Br</div>	<div>36 Kr</div>	<div>37 Rb</div>	<div>38 Sr</div>	<div>39 Y</div>	<div>40 Zr</div>	<div>41 Nb</div>	<div>42 Mo</div>	<div>43 Tc</div>	<div>44 Ru</div>	<div>45 Rh</div>	<div>46 Pd</div>	<div>47 Ag</div>	<div>48 Cd</div>	<div>49 In</div>	<div>50 Sn</div>	<div>51 Sb</div>	<div>52 Te</div>	<div>53 I</div>	<div>54 Xe</div>	<div>55 Cs</div>	<div>56 Ba</div>	<div>57 La</div>	<div>58 Ce</div>	<div>59 Pr</div>	<div>60 Nd</div>	<div>61 Pm</div>	<div>62 Sm</div>	<div>63 Eu</div>	<div>64 Gd</div>	<div>65 Tb</div>	<div>66 Dy</div>	<div>67 Ho</div>	<div>68 Er</div>	<div>69 Tm</div>	<div>70 Yb</div>	<div>71 Lu</div>	<div>72 Hf</div>	<div>73 Ta</div>	<div>74 W</div>	<div>75 Re</div>	<div>76 Os</div>	<div>77 Ir</div>	<div>78 Pt</div>	<div>79 Au</div>	<div>80 Hg</div>	<div>81 Tl</div>	<div>82 Pb</div>	<div>83 Bi</div>	<div>84 Po</div>	<div>85 At</div>	<div>86 Rn</div>
<div>57 La</div>	<div>58 Ce</div>	<div>59 Pr</div>	<div>60 Nd</div>	<div>61 Pm</div>	<div>62 Sm</div>	<div>63 Eu</div>	<div>64 Gd</div>	<div>65 Tb</div>	<div>66 Dy</div>	<div>67 Ho</div>	<div>68 Er</div>	<div>69 Tm</div>	<div>70 Yb</div>	<div>71 Lu</div>	<div>72 Hf</div>	<div>73 Ta</div>	<div>74 W</div>	<div>75 Re</div>	<div>76 Os</div>	<div>77 Ir</div>	<div>78 Pt</div>	<div>79 Au</div>	<div>80 Hg</div>	<div>81 Tl</div>	<div>82 Pb</div>	<div>83 Bi</div>	<div>84 Po</div>	<div>85 At</div>	<div>86 Rn</div>																																																															
<div>87 Fr</div>	<div>88 Ra</div>	<div>89 Ac</div>	<div>90 Th</div>	<div>91 Pa</div>	<div>92 U</div>	<div>93 Np</div>	<div>94 Pu</div>	<div>95 Am</div>	<div>96 Cm</div>	<div>97 Bk</div>	<div>98 Cf</div>	<div>99 Es</div>	<div>100 Fm</div>	<div>101 Md</div>	<div>102 No</div>	<div>103 Lr</div>	<div>104 Rf</div>	<div>105 Db</div>	<div>106 Sg</div>	<div>107 Bh</div>	<div>108 Hs</div>	<div>109 Mt</div>	<div>110 Ds</div>	<div>111 Rg</div>	<div>112 Cn</div>	<div>113 Nh</div>	<div>114 Fl</div>	<div>115 Mc</div>	<div>116 Lv</div>	<div>117 Ts</div>	<div>118 Og</div>																																																													

🔍 Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🧪

💡 Matter:

Anything that has **mass** and **occupies space** 🧪🌍

Examples: air, gases, liquids, solids!

🔬 Atoms:

Atoms are the **smallest, indivisible particle** 🧪👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

📌 Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Chemistry Study Guide Answers Content Mastery

Clare T. Furse



Chemistry Study Guide Answers Content Mastery:

Chemistry: Matter & Change, Study Guide For Content Mastery, Student Edition McGraw Hill, 2001-03-30 Study Guide and Reinforcement Worksheets allow for differentiated instruction through a wide range of question formats There are worksheets and study tools for each section of the text that help teachers track students progress toward understanding concepts Guided Reading Activities help students identify and comprehend the important information in each chapter

Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook McGraw Hill, 2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change Chemistry, 2001-04-01 **Chemistry** McGraw-Hill Staff, 2001-03 *Study Guide with Selected Solutions* Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems **Excel Science Study Guide, Years**

9-10 Will Marchment, 2004 The book contains coverage of five major topic areas in the NSW School Certificate test Energy Force and Motion Atoms Elements and Compounds Structure and Function of Living Things Earth and Space Ecosystems Resources and Technology a chapter on Investigations and Problem Solving in Science to help with practical skills revision questions and chapter tests to help you remember important information a glossary and summary in each section of the book diagrams and illustrations to help your understanding a section to help you prepare for the School Certificate test a sample School Certificate test paper with answers answers to all questions Chemistry - Study Guide to Accompany

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chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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