

Phase Change

- Deposition-gas to a solid; Sublimation-solid to gas; Vaporization- liquid to gas; Condensation-gas to liquid
- Physical change - melting, freezing, dissolving, evaporating;
- Chemical change- boiling an egg, converting, exploding, burning, digesting.

Atomic Numbers

- Number that identifies an element-Atomic number
- Weighted average mass of all occurring isotopes of an element-Atomic mass
- Symbolic method for expressing the composition of an atomic nucleus-Atomic notation

Electron Configurations

- Shorthand description of electrons by sublevels according to increasing energy-Electron configuration
- Sublevels: 2,6,10,14 electrons in the sublevels; 1s 2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d 7p;



Chemical Formulas

- Name of a chemical compound that indicates the # of atoms of each element-Chemical formula
- Number of atoms from formula- $\text{C}_7\text{H}_{16}\text{NO}_3$ = 40 atoms

NaHCO_3 -sodium hydrogen carbonate; $\text{Mg}(\text{ClO}_4)_2$ - Magnesium Perchlorate; PbO_2 - Lead (IV) oxide;
 Hg_2Br_2 -Mercury (II) bromide; $\text{Al}_2(\text{CrO}_4)_3$ -Aluminum Cromate; ZnCl_2 - Zinc Chloride; Nitrous acid- HNO_2 ;
Hypochlorous acid- HClO ; hydrochloric- HCl ; H_2S -hydrosulfuric acid; HClO_4 -perchloric acid; H_2SO_4 -sulfuric acid; HNO_3 -nitric acid; SnO_2 -tin dioxide; sulfuric acid- H_2SO_4 ;

Chemistry Study Guide Answers Phase And Changes

James E. Brady, John R. Holum

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Fundamentals of Chemistry, Study Guide James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs Student Study Guide Fundamentals of Chemistry by David E. Goldberg Margaret G. Kimble, 1994

Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter, 1993

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Prentice Hall Chemistry, 2000 **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999

Resources for Teaching Middle School Science Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-04-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them Resources for Teaching Middle School Science developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of Resources for Teaching Elementary School Science the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum materials included in this book were selected by panels of teachers and scientists using

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The National Guide to Educational Credit for Training Programs, 1999 **General Chemistry**

James E. Brady, 1986 **CHEMISTRY** SILBERBERG, 2003 *Report* National Research Council of Canada, 1918

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