

## Phase Change

- **Deposition**-gas to a solid; **Sublimation**-solid to gas; **Vaporization**- liquid to gas; **Condensation**-gas to liquid
- **Physical change** = melting, freezing, dissolving, evaporating;
- **Chemical change** = boiling an egg, converting, exploding, burning, digesting.

## Atomic Numbers

- Number that identifies an element-**Atomic number**
- Weighted average mass of all occurring isotopes of an element-**Atomic mass**
- Symbolic method for expressing the composition of an atomic nucleus-**Atomic notation**

## Electron Configurations

- Shorthand description of electrons by sublevels according to increasing energy-**Electron configuration**
- **Sublevels:** 2,6,10,14 electrons in the sublevels; **1s 2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d 7p;**



## Chemical Formulas

- Name of a chemical compound that indicates the # of atoms of each element-**Chemical formula**
- **Number of atoms** from formula- $\rightarrow \text{C}_{17}\text{H}_{19}\text{NO}_3 = 40\text{atoms}$

**NaHCO<sub>3</sub>**=sodium hydrogen carbonate; **Mg(ClO<sub>4</sub>)<sub>2</sub>**= Magnesium Perchlorate; **PbO<sub>2</sub>**=Lead (IV) oxide;

**Hg<sub>2</sub>Br<sub>2</sub>**=Mercury (II) bromide; **Al<sub>2</sub>(CrO<sub>4</sub>)<sub>3</sub>**=Aluminum Chromate; **ZnCl<sub>2</sub>**= Zinc Chloride; Nitrous acid=**HNO<sub>2</sub>**;

Hypochlorous acid=**HClO**; hydrochloric=**HCl**; **H<sub>2</sub>S**=hydrosulfuric acid; **HClO<sub>4</sub>**=perchloric acid; **H<sub>2</sub>SO<sub>4</sub>**=sulfuric acid; **HNO<sub>3</sub>**=nitric acid; **SnO<sub>2</sub>**=Tin dioxide; sulfuric acid=**H<sub>2</sub>SO<sub>4</sub>**;

# Chemistry Study Guide Answers Phase And Changes

**James E. Brady, John R. Holum**



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