

PROBLEMS AND QUESTIONS

1. Nitrogen gas and hydrogen gas combine according to the balanced equation below:



The diagram below shows an initial mixture of nitrogen gas (blue spheres) and hydrogen gas (white spheres) reacting to form ammonia gas (white spheres). Before the synthesis of the final product, ammonia, explain what is happening.



Initial mixture of gases



Final mixture of gases

Explanation:

There are 4 nitrogen molecules in the mixture. Each one is made up of 2 nitrogen atoms. If N_2 had 10 molecules, it would be 20 nitrogen atoms. There would be 6 hydrogen molecules. Each one is made up of 2 hydrogen atoms. If there were 10 molecules, it would be 20 hydrogen atoms. If there were 10 molecules, it would be 20 hydrogen atoms.

2. Use the same equation as question 1 to answer the following questions.

- a. If 28.0 g of N_2 and 4.0 g of H_2 are reacted, what mass of ammonia can be produced?

$$\begin{aligned} \text{N}_2 + 3\text{H}_2 &\rightarrow 2\text{NH}_3 \\ 28.0 \text{ g N}_2 &\times \frac{1 \text{ mol N}_2}{28.0 \text{ g N}_2} \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \times \frac{17.0 \text{ g NH}_3}{1 \text{ mol NH}_3} = 34.0 \text{ g NH}_3 \\ 4.0 \text{ g H}_2 &\times \frac{1 \text{ mol H}_2}{2.0 \text{ g H}_2} \times \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \times \frac{17.0 \text{ g NH}_3}{1 \text{ mol NH}_3} = 23.3 \text{ g NH}_3 \end{aligned}$$

- b. What is the limiting reagent? Explain your work.

Answer: H_2 is the limiting reagent.

Answer: H_2

Answer: H_2

- c. What is the percent yield if 34.0 g of ammonia is produced in this reaction?

$$\frac{34.0 \text{ g}}{34.0 \text{ g}} \times 100\%$$

Answer: 100%

Ccsf Chem 40 Lab Manual

Clark E. Bricker



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