

## Moles & Stoichiometry Cheat Sheet

<b>Calculating Molar Mass</b> <ol style="list-style-type: none"> <li>1. Write out formula of compound</li> <li>2. Determine number of atoms of each element present</li> <li>3. multiply number of atoms of the element by the atomic mass of the element</li> <li>4. add up the values just determined for all of the elements present in the compound</li> </ol>	<b>Calculating Percent Composition</b> <ol style="list-style-type: none"> <li>1. Calculate the molar mass of the compound</li> <li>2. divide the mass contribution of each element by the total molar mass</li> <li>3. multiply by 100 to make a percent</li> </ol> <p><i>Note: total of percents should equal 100%!</i></p>
<p><b>Mass (grams)</b></p> <p style="text-align: center;"> <math>\times</math> by molar mass  <math>\leftarrow \text{-----}</math>  <math>\text{-----} \rightarrow</math>  <math>\div</math> by molar mass         </p>	<p style="text-align: center;"><b>Mole</b></p> <p style="text-align: center;"> <math>\times</math> by Avogadro's #  <math>\text{-----}</math>  <math>\leftarrow \text{-----}</math>  <math>\text{-----} \rightarrow</math>  <math>\div</math> by Avogadro's #         </p> <p style="text-align: right;"><b>Particles</b> (atoms, molecules, formula units)</p>
<b>Calculating Empirical Formula</b> <ol style="list-style-type: none"> <li>1. Change % of each element to grams of each element (if no sample amount is given assume 100 g)</li> <li>2. convert grams of each element to moles of each element by dividing the grams by the molar (atomic) mass of that element</li> <li>3. divide all molar values by the smallest molar value (to establish lowest whole number ratio)</li> <li>4. if all values are not whole numbers, multiply all by a whole number factor to make the values whole (ex. 1.5 and 1 should be multiplied by 2 to get rid of the .5 = 3 and 2)</li> <li>5. plug values in with the element they represent</li> </ol>	<b>Calculating Molecular Formula</b> <ol style="list-style-type: none"> <li>1. calculate the empirical formula first</li> <li>2. calculate the molar mass of the empirical formula</li> <li>3. take the molar mass of the molecular formula (given in the problem) and divide it by the molar mass of the empirical formula. This will equal the number of times the e.f. is in the m.f.</li> <li>4. distribute the number attained through the empirical formula</li> </ol> <p style="text-align: center;"><i>Example: <math>2(\text{C}_2\text{H}_5) = \text{C}_4\text{H}_{10}</math></i></p>
<b>Calculating the Formula of a Hydrate</b> <ol style="list-style-type: none"> <li>1. Determine the mass of water in the hydrate by subtracting the mass of the anhydrous from the mass of the hydrate (if necessary) -OR- convert %'s into grams by assuming a 100g sample</li> <li>2. convert grams of each into moles of each by dividing by molar mass</li> <li>3. divide the moles of water by the moles of the anhydrous to determine the number of moles of water in the formula</li> <li>4. write the formula of the anhydrous followed by a dot then the # of moles of water and "<math>\text{H}_2\text{O}</math>"</li> </ol> <p style="text-align: center;"><i>Example: <math>\text{CuSO}_4 \cdot 5\text{H}_2\text{O}</math></i></p>	<b>Naming a Hydrate</b> <ol style="list-style-type: none"> <li>1. write the name of the ionic compound</li> <li>2. use a prefix for the number of water molecules (see below)</li> <li>3. write "hydrate" after the prefix</li> </ol> <p>1 = mono, 2 = di, 3 = tri, 4 = tetra, 5 = penta,          6 = hexa, 7 = hepta, 8 = octa, 9 = nona, 10 = deca</p> <p style="text-align: center;"><i>Example: Copper (II) sulfate pentahydrate</i></p>

# Cheat Sheet For Stoichiometry

**Sunil Tanna**



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Stoichiometry Sydney Young, 1918 This introduction was originally prefixed to Dr Findlay s Phase rule which

was the first volume of the series issued It belongs properly however to this volume and is therefore included here Note p vii Introduction issued also separately in 1904

**You Can Do Chemistry** Sunil Tanna,2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

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**Formulation and Stoichiometry** Emil J. Margolis,2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction

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web krzysztof wodiczek is professor in residence of art design and the public domain emeritus at the graduate school of design at harvard university he is renowned for his large scale slide and video projections on architectural facades and monuments

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web krzysztof wodiczek is a former director of the center for advanced visual studies and the head of interrogative design group at mit and since 2010 is a professor and a coordinator of art design and the public domain a postgraduate mdes concentration at the graduate school of design at harvard university

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