

(a) Write the chemical equation of the reaction which takes place on reacting  $\text{XNO}_3$  solution and common salt solution giving the physical states of all the reactants and products.  
 (b) What type of chemical reaction is illustrated by the above equation?

64. Two metals X and Y form the salts  $\text{XSO}_4$  and  $\text{Y}_2\text{SO}_4$ , respectively. The solution of salt  $\text{XSO}_4$  is blue in colour whereas that of  $\text{Y}_2\text{SO}_4$  is colourless. When barium chloride solution is added to  $\text{XSO}_4$  solution, then a white precipitate Z is formed alongwith a salt which turns the solution green. And when barium chloride solution is added to  $\text{Y}_2\text{SO}_4$  solution, then the same white precipitate Z is formed alongwith colourless common salt solution.  
 (a) What could the metals X and Y be?  
 (b) Write the name and formula of salt  $\text{XSO}_4$ .  
 (c) Write the name and formula of salt  $\text{Y}_2\text{SO}_4$ .  
 (d) What is the name and formula of white precipitate Z?  
 (e) Write the name and formula of the salt which turns the solution green in the first case.

65. A red-brown metal X forms a salt  $\text{XSO}_4$ . When hydrogen sulphide gas is passed through an aqueous solution of  $\text{XNO}_3$ , then a black precipitate of  $\text{X}_2\text{S}$  is formed alongwith sulphuric acid solution.  
 (a) What could the salt  $\text{XNO}_3$  be?  
 (b) What is the colour of salt  $\text{XNO}_3$ ?  
 (c) Name the black precipitate  $\text{X}_2\text{S}$ .  
 (d) By using the formula of the salt obtained in (a) above, write an equation of the reaction which takes place when hydrogen sulphide gas is passed through the aqueous solution.  
 (e) What type of chemical reaction takes place in this case?

66. When a strip of red-brown metal X is placed in a colourless salt solution  $\text{YNO}_3$ , then metal Y is set free and a blue coloured salt solution  $\text{X}_2\text{NO}_3$  is formed. The liberated metal Y forms a shining white deposit on the strip of metal X.  
 (a) What do you think metal X is?  
 (b) Name the salt  $\text{YNO}_3$ .  
 (c) What could be metal Y?  
 (d) Name the salt  $\text{X}_2\text{NO}_3$ .  
 (e) What type of reaction takes place between metal X and salt solution  $\text{YNO}_3$ ?

67. A metal salt MX when exposed to light splits up to form metal M and a gas  $\text{X}_2$ . Metal M is used in making ornaments whereas gas  $\text{X}_2$  is used in making bleaching powder. The salt MX is itself used in black and white photography.  
 (a) What do you think metal M is?  
 (b) What could be gas  $\text{X}_2$ ?  
 (c) Name the metal salt MX.  
 (d) Name any two salt solutions which on mixing together can produce a precipitate of salt MX.  
 (e) What type of chemical reaction takes place when salt MX is exposed to light? Write the equation of the reaction.

## ANSWERS

1. Decomposition reaction      4.  $\text{Zn} + 2\text{AgNO}_3 \longrightarrow \text{Zn}(\text{NO}_3)_2 + 2\text{Ag}$       8. Rancidity      6. Anti-oxidants  
 11. Decomposition reactions carried out by electricity      14.  $2\text{AgNO}_3 \text{ (aq)} + \text{Cu} \text{ (s)} \longrightarrow \text{Cu}(\text{NO}_3)_2 \text{ (aq)} + 2\text{Ag} \text{ (s)}$       15. (i) Decomposition (ii) Combination      16. (i) Decomposition (ii) Decomposition  
 (iii) Decomposition      17. (i) Displacement reaction (ii) Combination reaction      18. (a) See page 29 (b) See page 29      19. (a) Displacement reaction      (b) Combination reaction      20. (a) Combination reaction  
 (b) Displacement reaction      (c) Displacement reaction      (d) Decomposition reaction      (e) Double displacement reaction      21. (a)  $\text{H}_2\text{O}_2$  (b)  $\text{PbS}$       22.  $\text{H}_2\text{S}$       23. Substance oxidised :  $\text{H}_2\text{S}$  ; Substance reduced :  $\text{SO}_2$   
 24. (a) oxidation ; reduction (b) reduction ; oxidation (c) rancidity      25. (i) C (ii)  $\text{ZnO}$       26. (a) Magnesium is oxidised ; Addition of oxygen to magnesium takes place      (i) (i)  $\text{HCl}$  (ii)  $\text{MnO}_2$  (iii)  $\text{MnO}_3$  (iv)  $\text{HCl}$   
 27.  $\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}$  (i) Zinc oxide (ii) Carbon      28. (i)  $2\text{Cu} + \text{O}_2 \longrightarrow 2\text{CuO}$  (ii)  $\text{O}_2$   
 29. (i)  $2\text{Mg} + \text{O}_2 \longrightarrow 2\text{MgO}$  (ii) Mg (iii)  $\text{O}_2$       37. (i)  $\text{Cu} \text{ (s)} + 2\text{AgNO}_3 \text{ (aq)} \longrightarrow \text{Cu}(\text{NO}_3)_2 \text{ (aq)} + 2\text{Ag} \text{ (s)}$       39. (a)  $\text{H}_2$  (b)  $\text{CuO}$  (c)  $\text{H}_2$       43. (b) Rusting      (c) Oxidation      44. (b) Oxidation

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**L Cohen**

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