

10**CHEMICAL QUANTITIES****Practice Problems**

On a piece of paper, solve the following problems.

YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

1. What is the molar mass of sucrose ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$)?
2. What is the molar mass of each of the following compounds?
 - a. phosphorus pentachloride (PCl_5)
 - b. uranium hexafluoride (UF_6)
3. Calculate the molar mass of each of the following ionic compounds:
 - a. KMnO_4
 - b. $\text{Ca}_3(\text{PO}_4)_2$
4. How many moles is 3.52×10^{24} molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen (O_2)?

SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS

1. What is the molar mass of each of the following compounds?
 - a. $\text{C}_6\text{H}_{12}\text{O}_6$
 - b. NaHCO_3
 - c. C_3H_8
 - d. KNH_4SO_4
2. Calculate the mass in grams of each of the following:
 - a. 8.0 mol lead oxide (PbO)
 - b. 0.75 mol hydrogen sulfide (H_2S)
 - c. 0.00100 mol silicon tetrahydride (SiH_4)
 - d. 1.50×10^{-2} mol molecular oxygen (O_2)
 - e. 2.30 mol ethylene glycol ($\text{C}_2\text{H}_6\text{O}_2$)
3. How many grams are in 1.73 mol of dinitrogen pentoxide (N_2O_5)?
4. How many grams are in 0.658 mol of calcium phosphate [$\text{Ca}_3(\text{PO}_4)_2$]?
5. Calculate the number of moles in each of the following:
 - a. 0.50 g sodium bromide (NaBr)
 - b. 13.5 g magnesium nitrate [$\text{Mg}(\text{NO}_3)_2$]
 - c. 1.02 g magnesium chloride (MgCl_2)
 - d. 0.00100 g monochloromethane (CH_3Cl)
 - e. 1.50×10^{-3} g propylene glycol [$\text{C}_3\text{H}_8(\text{OH})_2$]
6. A chemist plans to use 435.0 grams of ammonium nitrate (NH_4NO_3) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ($\text{C}_{20}\text{H}_{24}\text{N}_2\text{O}_2$). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane (CH_4) gas?
9. How many moles is 135 L of ammonia (NH_3) gas at STP?

Chapter 10 Chemical Quantities Study Guide Answer Key

Rachel Sandford



Chapter 10 Chemical Quantities Study Guide Answer Key:

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The book delves into Chapter 10 Chemical Quantities Study Guide Answer Key. Chapter 10 Chemical Quantities Study Guide Answer Key is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Chapter 10 Chemical Quantities Study Guide Answer Key, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Chapter 10 Chemical Quantities Study Guide Answer Key
 - Chapter 2: Essential Elements of Chapter 10 Chemical Quantities Study Guide Answer Key
 - Chapter 3: Chapter 10 Chemical Quantities Study Guide Answer Key in Everyday Life
 - Chapter 4: Chapter 10 Chemical Quantities Study Guide Answer Key in Specific Contexts
 - Chapter 5: Conclusion
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 3. In chapter 2, this book will delve into the foundational concepts of Chapter 10 Chemical Quantities Study Guide Answer Key. This chapter will elucidate the essential principles that need to be understood to grasp Chapter 10 Chemical Quantities Study Guide Answer Key in its entirety.
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 6. In chapter 5, this book will draw a conclusion about Chapter 10 Chemical Quantities Study Guide Answer Key. The final chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Chapter 10 Chemical Quantities Study Guide Answer Key.

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