

**Chapter 10 Study Guide: The Mole****Matching:**

Match each term with the correct definition below.

- a. molar volume
- b. molar mass
- c. molar mass

- 1. the number of grams of an element that is numerically equal to the atomic mass of the element in amu
- 2. the mass of a mole of any element, or compound
- 3. the volume occupied by a mole of any gas at STP

Match each term with the correct definition below.

- a. representative particle
- b. mole
- c. Avogadro's number

- d. percent composition
- e. standard temperature and pressure
- f. empirical formula

- 4. the number of representative particles in 1 mole of that substance
- 5. an atom, an ion, or a molecule, depending upon the way a substance numerically exists
- 6. the SI unit used to measure amount of substance
- 7. EMC and 1 atom
- 8. the percent by mass of each element in a compound
- 9. the smallest whole number ratio of the atoms in a compound

**Multiple Choice:**

Identify the choice that best completes the statement or answers the question.

- 10. What SI unit is used to measure the number of representative particles in a substance?
  - a. kilogram
  - b. amper
  - c. liters
  - d. mole
- 11. How many hydrogen atoms are in 5 molecules of Isopropyl alcohol,  $C_3H_7OH$ ?
  - a.  $5 \times (6.02 \times 10^{23})$
  - b. 5
  - c.  $2.5 \times (6.02 \times 10^{23})$
- 12. Avogadro's number of representative particles is equal to one \_\_\_\_\_.
  - a. kilogram
  - b. gram
  - c. liters
  - d. mole

## Chapter 10 Study Guide The Mole Answer Key

**Alfred Marlyn Elliott, Robert Edward  
Hayes, John Ewald Lutz**

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