

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent	solvent
insoluble	solid	solid	solid

After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is the (3) _____ solvent _____ and solute is the

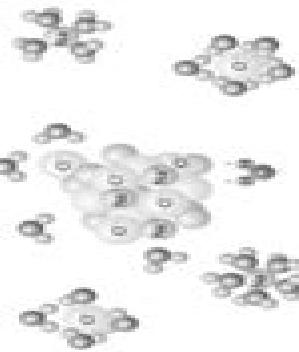
(4) _____ solution. A solution that does not dissolve is
 (5) _____ insoluble. (6) _____ liquid _____ solution is the most common
 type of solution. If one liquid is soluble in another liquid, such as water and oil, the
 two liquids are (7) _____ miscible. However, if one liquid is insoluble in another
 the liquids are (8) _____ immiscible.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Use the diagram to answer the following questions.



1. Indicate in solute is when the solute is in a solution. What is solute?

Solute is the process of surrounding solute particles with solvent particles to form a solution.

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CHAPTER 15 STUDY GUIDE FOR CONSTITUTANT MASTERY**Section 15.1** *Answers*

16. An sodium chloride dissolved in water. What happens is the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Unlike the attraction of the water molecule around the sodium ion and chloride ion.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom), because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a glass cup, add the rest of solute may be dissolved.

Starting at the bottom of the solution, breaking the solution into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, with type of solute.

True 14. The overall energy change that occurs when a solute dissolves is called the heat of solution.

False 15. Solubility is a property of the solute, amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

False 16. Solvent molecules in liquid as the solute has a low heat of solution.

False 17. In a saturated solution, solute and solvents are in equilibrium.

True 18. Additional solute can be dissolved in an unsaturated solution.

True 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Answer Key

Wayne S. Wooden, Randy Blazak

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Table of Contents Chapter 15 Study Guide For Content Mastery Solutions Answer Key

1. Understanding the eBook Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - The Rise of Digital Reading Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Personalized Recommendations
 - Chapter 15 Study Guide For Content Mastery Solutions Answer Key User Reviews and Ratings
 - Chapter 15 Study Guide For Content Mastery Solutions Answer Key and Bestseller Lists

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5. Accessing Chapter 15 Study Guide For Content Mastery Solutions Answer Key Free and Paid eBooks
 - Chapter 15 Study Guide For Content Mastery Solutions Answer Key Public Domain eBooks
 - Chapter 15 Study Guide For Content Mastery Solutions Answer Key eBook Subscription Services
 - Chapter 15 Study Guide For Content Mastery Solutions Answer Key Budget-Friendly Options
6. Navigating Chapter 15 Study Guide For Content Mastery Solutions Answer Key eBook Formats
 - ePUB, PDF, MOBI, and More
 - Chapter 15 Study Guide For Content Mastery Solutions Answer Key Compatibility with Devices
 - Chapter 15 Study Guide For Content Mastery Solutions Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Highlighting and Note-Taking Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Interactive Elements Chapter 15 Study Guide For Content Mastery Solutions Answer Key
8. Staying Engaged with Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 15 Study Guide For Content Mastery Solutions Answer Key
9. Balancing eBooks and Physical Books Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 15 Study Guide For Content Mastery Solutions Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Setting Reading Goals Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Fact-Checking eBook Content of Chapter 15 Study Guide For Content Mastery Solutions Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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