

## CHAPTER 19 STUDY GUIDE FOR CONTENT MASTERY

**Acids and Bases****Section 19.1 Acids and Bases: An Introduction**

In your notebook, read about the properties of acids and bases.

For each description below, write and *b* if the statement is a property of an acid or *b* if it is not a property of acids. If the property does not apply to either an acid or a base, write *n*.*Both* *b* Can an item from paper a different color?*Acid* *b* Does it turn指示剂 acidic?*Both* *b* Causes more hydrogen ions than hydronium ions.*Base* *b* Has a bitter flavor?*Both* *b* Does not conduct electricity.*Neither* *b* Has a high pH.*Neither* *b* Causes equal numbers of hydrogen and hydroxide ions.*Base* *b* Turns blue litmus.*Both* *b* Turns red litmus.

In your notebook, read about the different models of acids and bases.

Acid/base	Ionization theory	Octet rule and hydronium
Acidic base	hydrogen	octet rule and hydronium

Use the terms below to complete the sentence. You may use each term more than once.

The *(19)* \_\_\_\_\_ model of acids and bases states that an acid containsa base. A base contains the *(20)* \_\_\_\_\_ hydronium \_\_\_\_\_ group and electrons to prevent*(21)* \_\_\_\_\_ hydronium ions in aqueous solutions.According to the *(19)* \_\_\_\_\_ model, *(20)* \_\_\_\_\_ acidic and bases*(19)* \_\_\_\_\_ hydrogen ions, and a base accepts *(20)* \_\_\_\_\_ hydrogen ions.

According to the model, is an acid base, is acidic, and is basic?

*(19)* \_\_\_\_\_ conjugate base \_\_\_\_\_ and each has a *(19)* \_\_\_\_\_ conjugate acid \_\_\_\_\_.

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**CHAPTER 19 STUDY GUIDE FOR CONTENT MASTERY****Section 19.2 Strengths of Acids and Bases**

In your notebook, read about the strengths of acids.

Circle the letter of the choice that best completes the statement or answers the question.

1. Acid A and Acid B are of equal concentrations and are mixed with a conductivity apparatus. When the electrodes are placed in acid A, the bulb glows more brightly. When they are placed in acid B, the bulb glows more brightly. Which of the following is true?
  - a. Acid A is stronger than acid B.
  - b. Acid B is stronger than acid A.
  - c. Acid A and Acid B are of equal strength.
  - d. Strengths of strength can be made from the results.
2. A chemical equation for the ionization of an acid into a single ion in the right ( $\rightarrow$ ) is written for research and problem sets of the equation. Which of the following is true?
  - a. The anion does not indicate relative strength.
  - b. The anion and is half as strong.
  - c. The anion and is twice as strong.
  - d. The anion and is weak.
3. Strong acid is strong acid. What is true about its conjugate base?
  - a. Its conjugate base is explosive.
  - b. Its conjugate base is strong.
  - c. Its conjugate base is weak.
  - d. Its conjugate base is made by multiplying the strength of the conjugate base.
4. In addition, a weak acid produces
  - ① a mixture of anions and cations.
  - ② a mixture of anions and ions.
  - ③ anions, but no hydronium ions.
  - ④ cations, but no hydronium ions.
5. Why are *l* values of acid and anion?
  - a. The concentration of anions does not affect the ionization.
  - b. The equilibrium is not stable.
  - c. The solution contains a high concentration of anions and acid molecules.
  - d. The solution contains a high concentration of anions and acid molecules.
6. Which of the following describes exactly two weak acid and hydroxide ions in solution?
  - a. a strong acid
  - b. a strong base
  - c. a weak acid
  - d. a weak base
7. In general, conjugate bases have acidic and hydroxide ions are
  - a. strong acids
  - b. strong bases
  - c. weak acids
  - d. weak bases

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# Chapter 19 Study Guide Acids Bases Answers

**Linda Swisher, Gary A. Thibodeau**

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