

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent
insoluble	solid	solvent

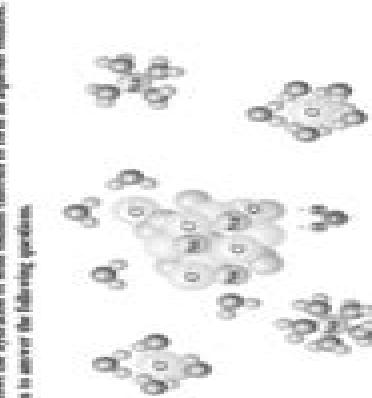
After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is dissolved in the (3) _____ solvent _____.

(4) _____ solute _____ is dissolved in a solvent, or (5) _____ is in (6) _____ solution. A substance that does not dissolve is (7) _____ insoluble _____ in (8) _____ liquid _____ solution or the most common type of solution. If one liquid is soluble in another liquid, such as water and oil, the two liquids are (9) _____ miscible _____ because if one liquid is insoluble in another, the liquids are (10) _____ immiscible _____.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.



The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Underline the diagram to answer the following questions.

1. In a solution, what is the solute? What is the solvent?
Solution is the process of **surrounding** solute particles with solvent particles to form a solution.

2. Explain the solubility of the solute in the solvent. _____

3. Explain the solubility of the solute in the solvent. _____

4. Explain the solubility of the solute in the solvent. _____

5. Explain the solubility of the solute in the solvent. _____

6. Explain the solubility of the solute in the solvent. _____

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CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Section 15.1** *Answers*

16. An sodium chloride dissolved in water, what happens is the sodium and chloride ions? The sodium and chloride ions are **separated** and surrounded by the water molecules.

17. Explain the solubility of the water molecule and the sodium ion and chloride ion. Because the sodium ion is **positively charged**, it attracts the **negatively charged** portion of the water molecule (the oxygen atom) and repels the **positively charged** portion of the water molecule (the hydrogen atom), because the chloride ion is **negatively charged**, it attracts the **positively charged** portion of the water molecule and repels the **negatively charged** portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How do you know? The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a glass cup, add the rest of solute may be dissolved, stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, write true or false.

True 16. The overall energy change that occurs when a solute dissolves is called the heat of solution.

False 17. Solubility is a property of the solute, amount of either the solute or a given amount of solute at a specified temperature and pressure.

False 18. Solute molecules in a glass of the solution are in contact with the glass walls.

True 19. In a saturated solution, solute and solvents are in equilibrium.

True 20. Additional solute can be dissolved in an unsaturated solution.

True 21. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Carolyn J. Meyer

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