

Chapter 18 - Reaction Rates and Equilibrium

Collision Theory

- Atoms, ions, and molecules can react to form products when they collide with one another
- Must have sufficient kinetic energy
- Lacking in kinetic energy, particles bounce apart unchanged
- Rate – a measure of speed of any change that occurs within an interval of time



Energy Changes in a Reaction

Activation energy –

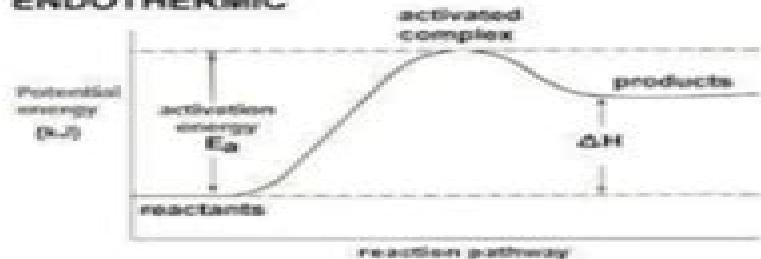
- minimum amount of KE that colliding particles need to react

Activated complex –

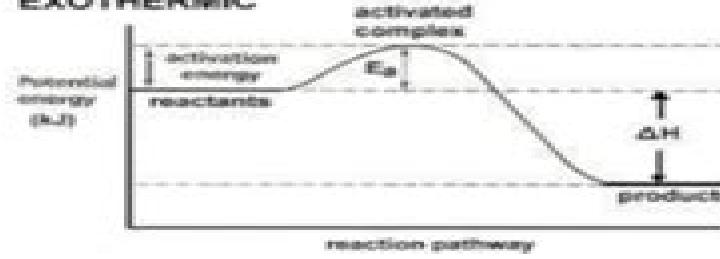
- temporary intermediate product
- may reform into the reactants or a new product

Potential Energy Diagram

ENDOTHERMIC



EXOTHERMIC



Factors Affecting Reaction Rates

1. Effect of Temperature

- Usually increasing temperature speeds up a reaction
- High kinetic energy

2. Effect of Concentration

- Increasing the number of particles into a fixed volume increases the collision frequency

3. Effect of Particle Size

- Increasing surface area increases the collision frequency

4. Effect of Catalysts

- Increases the rate of reaction without being used up itself in the reaction
- Lowers the activation energy



Chapter Assessment Reaction Rates Study Guide

Michael S. Matta

Chapter Assessment Reaction Rates Study Guide:

Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

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