

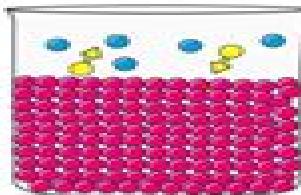
COLLIGATIVE PROPERTIES OF SOLUTIONS

A colligative property of a solution is a property that is only affected by the number of solute particles. There are four colligative properties of solutions, viz.:



Vapor Pressure Drop (ΔP)

When there is a lid on the container, the gas phase molecules are trapped; they are a vapor. The vapor creates a pressure.



If a solute is added to a solvent, the solute particles will prevent the solvent particles from turning into vapor. This causes a decrease in vapor pressure (ΔP).

Lots of dissolved substances $\rightarrow \Delta P$ big \rightarrow low vapor pressure (P).

The amount of vapor pressure drop can be determined by the following formula

a. Non-electrolyte solution

- $$\Delta P = P^* - P$$
$$\Delta P = X_t \times P^*$$
$$\Delta P = \frac{n_c}{n_c + n_p} \times P^*$$

- $$P = X_p \times P^*$$
$$P = \frac{n_p}{n_c + n_p} \times P^*$$

b. Electrolyte Solution

The electrolyte solution in the solvent can ionize. This results in the electrolyte solution having a more significant number of particles than the non-electrolyte solution at the same concentration. The degree of ionization is a direct measure of its ionizing state. The magnitude of the degree of ionization is expressed as :

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