

## COLLIGATIVE PROPERTIES OF SOLUTIONS

A colligative property of a solution is a property that is only affected by the number of solute particles. There are four colligative properties of solutions, viz :



### Vapor Pressure Drop ( $\Delta P$ )

When there is a lid on the container, the gas phase molecules are trapped, they are a vapor. The vapor creates a pressure



If a solute is added to a solvent, the solute particles will prevent the solvent particles from turning into vapor. This causes a decrease in vapor pressure ( $\Delta P$ ).

**Lots of dissolved substances  $\rightarrow \Delta P$  big  $\rightarrow$  low vapor pressure (P).**

The amount of vapor pressure drop can be determined by the following formula

#### a. Non-electrolyte solution

- $$\Delta P = P^* - P$$
$$\Delta P = X_t \times P^*$$
$$\Delta P = \frac{n_t}{n_t + n_p} \times P^*$$
- $$P = X_p \times P^*$$
$$P = \frac{n_p}{n_t + n_p} \times P^*$$

#### b. Electrolyte Solution

The electrolyte solution in the solvent can ionize. This results in the electrolyte solution having a more significant number of particles than the non-electrolyte solution at the same concentration. The degree of ionization is a direct measure of its ionizing state. The magnitude of the degree of ionization is expressed as :

# Colligative Properties Of Solutions Study Guide Answers

**Garrison**



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