

Chemistry Student Edition - Basic Answer Key

Chapter 11: Chemical Reactions

Chemical Equations

Questions

1. Early Chinese alchemists discovered an early form of gunpowder. What was the composition of this substance?
2. What later developments were made to the gunpowder recipe that improved its pyrotechnic properties?
3. Make an argument for why the burning of a candle is consistent with the law of conservation of matter/mass.
4. Think of an experiment that you could conduct to demonstrate that mass is conserved for a given chemical change.
5. Balance the following chemical equations:
 - a. $\text{C} + \text{O}_2 \rightarrow \text{CO}$
 - b. $\text{CO} + \text{O}_2 \rightarrow \text{CO}_2$
 - c. $\text{H}_2 + \text{Br}_2 \rightarrow \text{HBr}$
 - d. $\text{K} + \text{H}_2\text{O} \rightarrow \text{KOH} + \text{H}_2$
 - e. $\text{O}_3 \rightarrow \text{O}_2$
 - f. $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$
 - g. $\text{Zn} + \text{AgCl} \rightarrow \text{ZnCl}_2 + \text{Ag}$
 - h. $\text{Cl}_2 + \text{NaI} \rightarrow \text{NaCl} + \text{I}_2$
 - i. $\text{P}_4\text{O}_{10} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_4$
 - j. $\text{Be}_2\text{C} + \text{H}_2\text{O} \rightarrow \text{Be}(\text{OH})_2 + \text{CH}_4$
 - k. $\text{S} + \text{HNO}_3 \rightarrow \text{H}_2\text{SO}_4 + \text{NO}_2 + \text{H}_2\text{O}$
 - l. $\text{NH}_3 + \text{CuO} \rightarrow \text{Cu} + \text{N}_2 + \text{H}_2\text{O}$
 - m. $\text{HCl} + \text{CaCO}_3 \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$

Answers

1. Early gunpowder was comprised of saltpeter (impure potassium nitrate), sulfur, and charcoal (carbon).
2.
 - a. Purification of potassium nitrate
 - b. Using potassium chlorate instead of potassium nitrate
 - c. Further refinement of the recipe to find the proportions of reactants that yielded the most energetic reaction
3. The burning of a candle obeys the law of conservation of mass because, even though the mass of the candle will decrease as it burns, the mass is not lost. It is transferred into carbon dioxide and water that enter the surrounding air.
4. The reaction between iron and oxygen to produce rust could be studied as an example of conservation of mass. Steel wool could be placed into a covered beaker containing water. The reaction would proceed over time and rust would begin to form. The beaker could be weighed before and after the change. The mass would be the same, even though a new

Chapter 11 Chemical Reactions Answers

David R. Klein, Laurie S. Starkey



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Table of Contents Chapter 11 Chemical Reactions Answers

1. Understanding the eBook Chapter 11 Chemical Reactions Answers
 - The Rise of Digital Reading Chapter 11 Chemical Reactions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 11 Chemical Reactions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 11 Chemical Reactions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 11 Chemical Reactions Answers
 - Personalized Recommendations

- Chapter 11 Chemical Reactions Answers User Reviews and Ratings
- Chapter 11 Chemical Reactions Answers and Bestseller Lists
- 5. Accessing Chapter 11 Chemical Reactions Answers Free and Paid eBooks
 - Chapter 11 Chemical Reactions Answers Public Domain eBooks
 - Chapter 11 Chemical Reactions Answers eBook Subscription Services
 - Chapter 11 Chemical Reactions Answers Budget-Friendly Options
- 6. Navigating Chapter 11 Chemical Reactions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 11 Chemical Reactions Answers Compatibility with Devices
 - Chapter 11 Chemical Reactions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 11 Chemical Reactions Answers
 - Highlighting and Note-Taking Chapter 11 Chemical Reactions Answers
 - Interactive Elements Chapter 11 Chemical Reactions Answers
- 8. Staying Engaged with Chapter 11 Chemical Reactions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 11 Chemical Reactions Answers
- 9. Balancing eBooks and Physical Books Chapter 11 Chemical Reactions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 11 Chemical Reactions Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 11 Chemical Reactions Answers
 - Setting Reading Goals Chapter 11 Chemical Reactions Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 11 Chemical Reactions Answers
 - Fact-Checking eBook Content of Chapter 11 Chemical Reactions Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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