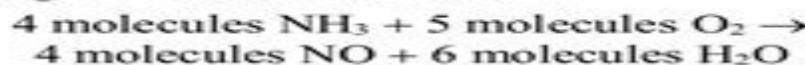


Chapter 11 (continued)



The coefficients represent both the numbers of particles and the numbers of moles interacting in the chemical reaction.



You can calculate the mass of each reactant and product by multiplying the number of moles by the conversion factor molar mass.

$$4 \cancel{\text{ mol NH}_3} \times \frac{17.03 \text{ g NH}_3}{1 \cancel{\text{ mol NH}_3}} = 68.12 \text{ g NH}_3$$

$$5 \cancel{\text{ mol O}_2} \times \frac{32.00 \text{ g O}_2}{1 \cancel{\text{ mol O}_2}} = 160.0 \text{ g O}_2$$

$$4 \cancel{\text{ mol NO}} \times \frac{30.01 \text{ g NO}}{1 \cancel{\text{ mol NO}}} = 120.0 \text{ g NO}$$

$$6 \cancel{\text{ mol H}_2\text{O}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \cancel{\text{ mol H}_2\text{O}}} = 108.1 \text{ g H}_2\text{O}$$

The law of conservation of mass is observed because the mass of the reactants ($68.12 \text{ g NH}_3 + 160.0 \text{ g O}_2 = 228.1 \text{ g}$) equals the mass of the products ($120.0 \text{ g NO} + 108.1 \text{ g H}_2\text{O} = 228.1 \text{ g}$).

Practice Problems

- Interpret each balanced equation in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.
 - $2\text{H}_2\text{O}_2(\text{l}) \rightarrow \text{O}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$
 - $\text{H}_2\text{CO}_3(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$
 - $4\text{HCl}(\text{aq}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + 2\text{Cl}_2(\text{g})$

As you know, the coefficients in a balanced chemical equation indicate the relationships among the moles of reactants and products in the reaction. You can use the coefficients to write mole ratios. A **mole ratio** is a ratio between the numbers of moles of any two substances in a balanced chemical equation. What mole ratios can be written for the following chemical equation?

Chapter 11 The Mole Study Guide Answer Key

**Barbara Hoffman, Robyn
Horsager, Scott Roberts, Vanessa
Rogers, Patricia Santiago-Muñoz, Kevin
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