
ANATOMY & PHYSIOLOGY

CHAPTER 2 STUDY GUIDE

TERM 1

Chemistry of Life

Levels of Chemical Organization

- **Atom:** The basic unit of matter, containing a nucleus made up of protons (positively charged) and neutrons (uncharged).
 - **Atomic Number:** Number of protons in the nucleus; defines the element.
 - **Atomic Mass:** Total number of protons and neutrons.
 - **Electrons:** Negatively charged particles that orbit the nucleus in energy levels (orbital), each capable of holding up to eight electrons.

Elements, Molecules, and Compounds

- **Element:** A pure substance made up of only one kind of atom.
- **Molecule:** A group of atoms bonded together.
- **Compound:** A substance with molecules that contain more than one kind of atom.

Chemical Bonding

- **Purpose:** Bonds allow stable atoms by filling their outer energy levels.
- **Types of Bonds:**
 - **Ionic Bonds:** Formed by the attraction of oppositely charged ions.
 - **Positive Ion:** Atom that loses electrons (e.g., Na^+ , Ca^{2+}).
 - **Negative Ion:** Atom that gains electrons (e.g., O^-).
 - **Covalent Bonds:** Formed by the sharing of electrons between atoms; generally stable and do not easily dissociate in water.
 - **Hydrogen Bonds:** Weak attractions that hold molecules together in polar charges or groups; do not cross-link molecules.

Chapter 12 Study Guide Blood

Linda Swisher, Gary A. Thibodeau

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Biology, the Science of Life, Third Edition W. J. Leverich, 1991 **Health: Teacher edition** Linda Brower Meeks, 1987

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- Chapter 4: Chapter 12 Study Guide Blood in Specific Contexts
- Chapter 5: Conclusion

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