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# ANATOMY & PHYSIOLOGY

## CHAPTER 2 STUDY GUIDE

WEEK 1

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### Chemistry of Life

#### Levels of Chemical Organization

- **Atom:** The basic unit of matter, containing a nucleus made up of protons (positively charged) and neutrons (uncharged).
    - **Atomic Number:** Number of protons in the nucleus; defines the element.
    - **Atomic Mass:** Total number of protons and neutrons.
    - **Electrons:** Negatively charged particles that orbit the nucleus in energy levels (orbitals), each capable of holding up to eight electrons.
- 

#### Elements, Molecules, and Compounds

- **Element:** A pure substance made up of only one kind of atom.
  - **Molecule:** A group of atoms bonded together.
  - **Compound:** A substance with molecules that contains more than one kind of atom.
- 

#### Chemical Bonding

- **Purpose:** Bonds form to stabilize atoms by filling their outer energy levels.
  - **Types of Bonds:**
    - **Ionic Bonds:** Formed by the attraction of oppositely charged ions.
      - **Positive Ion:** Atom that loses electrons (e.g.,  $\text{Na}^+$ ,  $\text{Ca}^{2+}$ ).
      - **Negative Ion:** Atom that gains electrons (e.g.,  $\text{Cl}^-$ ).
    - **Covalent Bonds:** Formed by the sharing of electrons between atoms; generally stable and do not easily dissociate in water.
    - **Hydrogen Bonds:** Weak attractions that hold molecules together in folded shapes or groups; do not create new molecules.
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# Chapter 12 Study Guide Blood

**Linda Swisher, Gary A. Thibodeau**



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