

General Chemistry II Chapter 13 Study Guide

- Raoult's Law: $P_{\text{sol}} = X_{\text{sol}} \cdot P_{\text{sol}^{\text{pure}}}$
 - P_{sol} = vapor pressure of solvent "a" above a solution
 - $P_{\text{sol}^{\text{pure}}}$ = vapor pressure of pure solvent "a"
 - X_{sol} = mole fraction of solvent "a" in the solution
 - $X_{\text{sol}} = \text{moles of solvent "a"} / \text{moles of solution (solvent particles + solute particles)}$
- Volatile: substances that easily evaporate and exhibit a vapor pressure
- Non-volatile: substances that have little tendency to evaporate
- Vapor pressure of the solution is the sum of the partial pressures of the solute and the solvent
 - If solute is non-volatile, vapor pressure of solution only comes from the solvent ($P_{\text{solution}} < P_{\text{pure solvent}}$)
 - If solute is more volatile than solvent, $P_{\text{solution}} > P_{\text{pure solvent}}$

Review on boiling point:

- Boiling point: temperature where vapor pressure = atmospheric pressure
 - Vapor pressure of liquid increases as temperature increases
 - Starts boiling process when vapor pressure = atmospheric pressure
 - Solute in solution lowers the vapor pressure of solvent- vapor pressure lowering
 - A higher temperature needed for vapor pressure to hit atmospheric pressure
 - boiling point of solution is higher than boiling point of pure solvent- boiling point elevation

Boiling point elevation:

- boiling point elevation: the boiling point of a solution is always higher than that of the pure solvent; it depends on the type of solvent and the # of solute particles
- $\Delta T_b = iK_b m$; $\Delta T_b = T_b - T_b^{\text{pure}}$
 - m = solute molality (moles solute / kg solvent)
 - K_b = molal boiling point elevation constant ($^{\circ}\text{C} / \text{m}$); depends only on solvent
 - i = van't Hoff factor
 - T_b = boiling point of solution
 - T_b^{pure} = boiling point of pure solvent

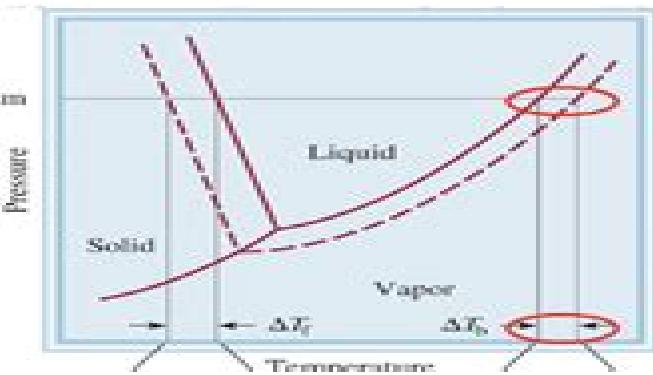


TABLE 13.2 Molal Boiling-Point Elevation and Freezing-Point Depression Constants of Several Common Liquids

Solvent	Normal Freezing Point ($^{\circ}\text{C}$)*	K_f ($^{\circ}\text{C}/\text{m}$)	Normal Boiling Point ($^{\circ}\text{C}$)*	K_b ($^{\circ}\text{C}/\text{m}$)
Water	0	1.86	100	0.52
Benzene	5.5	5.12	80.1	2.53
Ethanol	-117.3	1.99	78.4	1.22
Acetic acid	16.6	3.90	117.9	2.93
Cyclohexane	6.6	20.0	80.7	2.79

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John A. Olmsted, Gregory M. Williams

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