

# CHAPTER 15 STUDY GUIDE

## Energy and Chemical Change

### Section 15.1 Energy

*In your textbook, read about the nature of energy.*

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- \_\_\_\_\_ 1. *Energy* is the ability to do work or produce heat.
- \_\_\_\_\_ 2. The law of conservation of energy states that energy *can be* created and destroyed.
- \_\_\_\_\_ 3. Chemical potential energy is energy stored in a substance because of its *composition*.
- \_\_\_\_\_ 4. *Heat* is a form of energy that flows from a warmer object to a cooler object.
- \_\_\_\_\_ 5. A calorie is the amount of energy required to raise the temperature of *one gram* of pure water by one degree Celsius.
- \_\_\_\_\_ 6. A *calorie* is the SI unit of heat and energy.
- \_\_\_\_\_ 7. The *specific* heat of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius.
- \_\_\_\_\_ 8. *Kinetic energy* is energy of motion.
- \_\_\_\_\_ 9. Chemicals participating in a chemical reaction contain only *potential energy*.
- \_\_\_\_\_ 10. One nutritional Calorie is equal to *100 calories*.
- \_\_\_\_\_ 11. One calorie equals *4.184 joules*.
- \_\_\_\_\_ 12. When a fuel is burned, some of its *chemical potential energy* is lost as heat.
- \_\_\_\_\_ 13. To convert kilojoules to joules, *divide* the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

# Chapter 15 Chemistry Study Guide Answers

**T. W. Graham Solomons**



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**Wastewater Operator Certification Study Guide** John Giorgi, 2011-01-12      **Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e** Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry      **Study Guide to Accompany Basics for Chemistry** Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students      **Student Study Guide Fundamentals of Chemistry by David E. Goldberg** Margaret G. Kimble, 1994      Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines      **Study Guide, Answer Book to Accompany Organic Chemistry** Alan S. Wingrove, Robert L. Caret, 1981      Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed Robert K. Wismer, 1982      *General Chemistry Study Guide Sixth Edition* Darrell D. Ebbing, 1999      Organic

Chemistry, Study Guide and Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, 2000 On the cover of this book is a Pacific yew tree found in the ancient forests of the Pacific Northwest The bark of the Pacific yew tree produces Taxol found to be a highly effective drug against ovarian and breast cancer Taxol blocks mitosis during eukaryotic cell division The supply of Taxol from the Pacific yew tree is vanishingly small however A single 100 year old tree provides only about one dose of the drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It s just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

**Chemistry, Student Study Guide** John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that s real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You ll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

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**Organic Chemistry, Study Guide** T. W. Graham Solomons, 1996 This supplement accompanies an updated text which features more than 350 new problems Carbonyl chemistry is now covered in consecutive chapters The concept of ionic reactions is consolidated before students move on to

study radical reactions      *Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual* T. W. Graham Solomons, 1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry      CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book      *Chemistry, Study Guide* Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460      Study Guide to Chemistry and Life John William Hill, Dorothy M. Feigl, 1978      **Student Study Guide to Accompany Petrucci's General Chemistry** Robert K. Wismer, 1985      **Study Guide with Solutions to Accompany Burgoyne, A Study Course in Organic Chemistry** N. Sue Barnes, 1979      **Organic Chemistry Study Guide & Problems Book** Brent L. Iverson, 1998 Companion to Organic chemistry 2nd ed by William H Brown Provides a detailed section by section overview of the major points covered in the text All of the problems from the text are reprinted here with detailed stepwise solutions

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