

CHAPTER 19 STUDY GUIDE FOR CONTENT MASTERY

Acids and Bases**Section 19.1 Acids and Bases: An Introduction**

In your notebook, read about the properties of acids and bases.

For each description below, write and *b* to show a property of an acid or *b* to show a property of a base. If the property does not apply to either an acid or a base, write *neither*.**acid** *b* Can base form a hydrogen ion**acid** *b* Shows acidic taste**acid** *b* Releases more hydrogen ions than hydronium ions**base** *b* Has basic taste**base** *b* Has basic properties**neither** *b* Has neutral taste**neither** *b* Contains equal numbers of hydrogen and hydroxide ions**base** *b* Turns blue**base** *b* Turns red

In your notebook, read about the different models of acids and bases.

Use the following table to complete the passage. You may use each term more than once.

Acid/base	Ionization theory	Octet rule and hydroxide
acidic base	hydrogen	octet rule

The **(19)** **Arrhenius** model of acids and bases states that an acid containsan excess of **(20)** **hydrogen** ions and an acid of this character when it is dissolved in water. A base contains the **(21)** **hydroxide** group and dissociates to produce**(22)** **hydroxide** ions in aqueous solutions.According to the **(23)** **ionization theory**, acids and bases**(24)** **hydrogen** ions and base **(25)** **hydroxide** ions in aqueous

According to the model, in an aqueous solution, each acid is a

(26) **conjugate base** and each base is a **(27)** **conjugate acid**.

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Section 19.2 Strengths of Acids and Bases

In your notebook, read about the strengths of acids.

Circle the letter of the choice that best completes the statement or answers the question.

- Acid A and Acid B are of equal concentrations and are mixed with a conductivity apparatus. When the electrodes are placed in acid A, the bulb glows more brightly. When they are placed in acid B, the bulb glows more brightly. Which of the following is true?
 - a** Acid A is stronger than acid B.
 - b** Acid B is stronger than acid A.
 - c** Acid A and Acid B are of equal strength.
 - d** Strengths of strength can be read from the meter.
- A chemical equation for the ionization of an acid into a single ion in the right (\rightarrow) is written for research and problem sets of the chapter. Which of the following is true?
 - a** The answer does not indicate either strength. **③** The ionizing acid is strong.
 - b** The ionizing acid is half ionized. **④** The ionizing acid is weak.
 - c** The ionizing acid is strong. **②** The ionizing acid is weak.
 - d** Both the acid is strong and **①** What is true about the conjugate base?
- Acids and bases are **(28)** **electrolytes**.
 - a** In aqueous form is explosive. **③** In aqueous form produces
 - b** In aqueous form is strong. **④** In aqueous form is weak.
 - c** In aqueous form is weak. **②** In aqueous form is strong.
 - d** None of the above.
- For research sets to make comparing the strength of the conjugate base.
 - a** In aqueous form is explosive. **③** In aqueous form produces
 - b** In aqueous form is strong. **④** In aqueous form is weak.
 - c** In aqueous form is weak. **②** In aqueous form is strong.
 - d** None of the above.
- For research sets to make comparing the strength of the conjugate acid.
 - a** The concentration of water does not affect the ionization. **③** The concentration of water does affect the ionization.
 - b** The ionization is not acidic. **④** The ionization is not basic.
 - c** The solution contains a high concentration of ionized acid molecules. **②** The solution contains a high concentration of ionized acid molecules.
 - d** Both of the following directions identify the water acidic and hydroxide base is acidic.
 - ③** a strong acid **④** a weak acid
 - ④** a weak acid **③** a strong base
- In general, conjugate bases form from acids and hydroxide ions are
 - a** strong acids. **③** strong bases. **④** weak acids. **②** weak bases.
 - b** strong acids. **④** strong bases. **③** weak acids. **①** weak bases.
 - c** strong acids. **②** strong bases. **①** weak acids. **④** weak bases.
 - d** strong acids. **①** strong bases. **③** weak acids. **②** weak bases.

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Chapter 19 Acids And Bases Study Guide

David R. Klein

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