

CHAPTER 19 STUDY GUIDE FOR CONTENT MASTERY

Acids and Bases

Section 19.1 Acids and Bases: An Introduction

In your notebook, read about the properties of acids and bases.

For each description below, write acid if it tells about a property of an acid or base if it tells about a property of a base. If the property does not apply to either an acid or a base, write neither. If it applies to both an acid and a base, write both.

1. Can turn litmus paper a different color both
2. Reacts with certain metals acid
3. Contains more hydrogen ions than hydroxide ions acid
4. Feels slippery base
5. Reacts with carbonates acid
6. Feels rough neither
7. Contains equal numbers of hydrogen and hydroxide ions neither
8. Tastes bitter base
9. Tastes sour acid

In your notebook, read about the different models of acids and bases.

Use the terms below to complete the passages. You may use each term more than once.

Arrhenius conjugate base	Brønsted-Lowry hydrogen	conjugate acid hydroxide
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The (10) Arrhenius model of acids and bases states that an acid contains the element (11) hydrogen and forms ions of this element when it is dissolved in water. A base contains the (12) hydroxide group and dissociates to produce (13) hydroxide ions in aqueous solution.

According to the (14) Brønsted-Lowry model, an acid donates

(15) hydrogen ions, and a base accepts (16) hydrogen ions.

According to this model, in an acid-base reaction, each acid has a

(17) conjugate base and each base has a (18) conjugate acid.

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Section 19.2 Strengths of Acids and Bases

In your notebook, read about strengths of acids.

Circle the letter of the choice that best completes the statement or answers the question.

1. Acid A and acid B are of equal concentration and are tested with a conductivity apparatus. When the electrodes are placed in acid A, the bulb glows dimly. When they are placed in acid B, the bulb glows more brightly. Which of the following is true?
 - a. Acid A is stronger than acid B.
 - b. Acid B is stronger than acid A.
 - c. Acid A and acid B are of equal strength.
 - d. No comparison of strength can be made from the results.
2. A chemical equation for the ionization of an acid runs a single arrow to the right (\rightarrow) to separate the reactant and product sides of the equation. Which of the following is true?
 - a. The arrow does not indicate relative strength.
 - b. The ionizing acid is strong.
 - c. The ionizing acid is half ionized.
 - d. The ionizing acid is weak.
3. Sulfuric acid is a strong acid. What is true about its conjugate base?
 - a. Its conjugate base is amphiprotic.
 - b. Its conjugate base is strong.
 - c. Its conjugate base is weak.
 - d. No conclusion can be made regarding the strength of the conjugate base.
4. In solution, a weak acid produces
 - a. a mixture of molecules and ions.
 - b. all ions.
 - c. all molecules.
 - d. anions, but no hydroxide ions.
5. Why are K_a values of weak numbers?
 - a. The concentration of water does not affect the ionization.
 - b. The equilibrium is not stable.
 - c. The solution contains a high concentration of ions.
 - d. The solution contains a high concentration of un-ionized acid molecules.
6. Which of the following dissociates nearly one hundred percent in hydroxide ions in solution?
 - a. a strong acid
 - b. a strong base
 - c. a weak acid
 - d. a weak base
7. In general, compounds formed from active metals and hydroxide ions are
 - a. strong acids
 - b. strong bases
 - c. weak acids
 - d. weak bases

Chapter 19 Acids And Bases Study Guide

David R. Klein



Chapter 19 Acids And Bases Study Guide:

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