

**CHAPTER 9** STUDY GUIDE FOR CONTENT MASTERY

## Covalent Bonding

**Section 11 The Covalent Bond**

In your textbook, read about the nature of covalent bonds.

Underline the terms below just now to complete the passage.

covalent bond	valence	sigma bond	double bond	pi bond
When sharing of electrons occurs, the electrons between atoms that results is called				

(a) **covalent bond** \_\_\_\_\_ When such an electron is located between two atoms, it is shared, and the position is (b) **sigma bonds** \_\_\_\_\_ When two or more atoms bond by means of sharing, during the resulting patch is called an (c) **multiple bond** \_\_\_\_\_ If the electrons shared are located between the two atoms, the valence is called and (d) **sigma bond** \_\_\_\_\_ If the electrons shared are located between the two atoms, the pi bond is located between the two atoms and is (e) **pi bond** \_\_\_\_\_

In your textbook, read about single and double bonds and bond strength.

Circle the letter of the theory that best completes the statement to answer the question.

b. In what form do electrons in two hydrogen atoms, oxygen, and nitrogen normally move?

a. as single atoms

b. as molecules containing two atoms

c. as molecules containing four atoms

d. as molecules containing two atoms

e. as molecules containing the same substance

f. as molecules containing two different substances

g. as molecules containing two identical atoms

h. as molecules containing two different atoms

i. as molecules containing four different atoms

j. as molecules containing two different substances

k. as molecules containing two different substances

l. as molecules containing four different substances

m. as molecules containing two different substances

n. as molecules containing four different substances

o. as molecules containing two different substances

p. as molecules containing four different substances

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## Section 12 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write true or false.

(a) **True** \_\_\_\_\_ Every molecular compound is generally composed of a metal and a nonmetal.

(b) **True** \_\_\_\_\_ The general formula for the formula of a binary compound is  $M_nX_m$  where  $M$  is the symbol for the metal and  $X$  is the symbol for the nonmetal.

(c) **True** \_\_\_\_\_ The prefix *mono* indicates that there are two atoms.

(d) **True** \_\_\_\_\_ The prefix *di* indicates that there are two atoms.

(e) **True** \_\_\_\_\_ In naming the first element in a formula, the prefix *mono* is not used.

(f) **True** \_\_\_\_\_ For binary acids, the hydrogen part of the compound is named using the prefix *hydro*.

(g) **True** \_\_\_\_\_ A binary acid contains only two elements.

(h) **True** \_\_\_\_\_ If the name of the element in an oxyacid ends in *o*, the acid name contains the suffix *-ic*.

In your textbook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

**Column A**

(a) **Li** \_\_\_\_\_ **Li**  $Cl$

(b) **Na** \_\_\_\_\_  $Na$   $Cl$

(c) **Ca** \_\_\_\_\_  $Ca$   $Cl$

(d) **Al** \_\_\_\_\_  $Al$   $Cl$

(e) **Na** \_\_\_\_\_  $Na$   $NO_3$

(f) **Ca** \_\_\_\_\_  $Ca$   $NO_3$

(g) **Li** \_\_\_\_\_  $Li$   $NO_3$

(h) **Na** \_\_\_\_\_  $Na$   $NO_3$

(i) **Ca** \_\_\_\_\_  $Ca$   $NO_3$

(j) **Al** \_\_\_\_\_  $Al$   $NO_3$

(k) **Li** \_\_\_\_\_  $Li$   $NO_3$

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# Chapter 7 Study Guide For Content Mastery Answer Key

**Bernadette Mazurek Melnyk, Ellen  
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