

CONCEPTUAL Physics PRACTICE PAGE**Chapter 4 Newton's Second Law of Motion**
Mass and Weight

Learning physics is learning the connections among concepts in nature, and also learning to distinguish between closely-related concepts. Velocity and acceleration, previously treated, are often confused. Similarly in this chapter, we find that mass and weight are often confused. They aren't the same! Please review the distinction between mass and weight in your textbook. To reinforce your understanding of this distinction, circle the correct answers below.



Comparing the concepts of mass and weight, one is basic—fundamental—depending only on the internal makeup of an object and the number and kind of atoms that compose it. The concept that is fundamental is [mass] [weight].

The concept that additionally depends on location in a gravitational field is [mass] [weight].

[Mass] [Weight] is a measure of the amount of matter in an object and only depends on the number and kind of atoms that compose it.

It can correctly be said that [mass] [weight] is a measure of "laziness" of an object.

[Mass] [Weight] is related to the gravitational force acting on the object.

[Mass] [Weight] depends on an object's location, whereas [mass] [weight] does not.

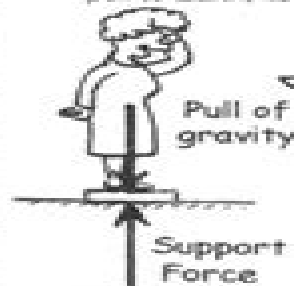
In other words, a stone would have the same [mass] [weight] whether it is on the surface of Earth or on the surface of the Moon. However, its [mass] [weight] depends on its location.

On the Moon's surface, where gravity is only about $1/6^{\text{th}}$ Earth gravity [mass] [weight]
[both the mass and the weight] of the stone would be the same as on Earth.

While mass and weight are not the same, they are [directly proportional] [inversely proportional] to each other. In the same location, twice the mass has [twice] [half] the weight.

The Standard International (SI) unit of mass is the [kilogram] [newton], and the SI unit of force is the [kilogram] [newton].

In the United States, it is common to measure the mass of something by measuring its gravitational pull to Earth, its weight. The common unit of weight in the U.S. is the [pound] [kilogram] [newton].



When I step on a weighing scale, two forces act on it; a downward pull of gravity, and an upward support force. These equal and opposite forces effectively compress a spring inside the scale that is calibrated to show weight. When in equilibrium, my weight = mg .

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web aug 19 2022 atoms with the same number of protons but different numbers of neutrons are called isotopes they share almost the same chemical properties but differ in mass and therefore in physical properties there are stable isotopes which do not emit radiation and there are unstable isotopes which do emit radiation

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web atoms of the same element with different numbers of neutrons are called isotopes close isotope atoms of an element with the same number of protons and electrons but different numbers of neutrons

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web oct 31 2023 an isotope is one of two or more species of atoms of a chemical element with the same atomic number and

position in the periodic table and nearly identical chemical behavior but with different atomic masses and physical properties every chemical element has one or more isotopes

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web isotopes are distinct nuclear species or nuclides as technical term of the same chemical element they have the same atomic number number of protons in their nuclei and position in the periodic table and hence belong to the same chemical element but differ in nucleon numbers mass numbers due to different numbers of neutrons in

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web verified by toppr correct option is d atoms having same atomic number but different atomic mass are known as isotopes the difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons

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web part i answer the questions based on the above reading what is an isotope isotopes are versions of the same element they have the same number of protons and electrons as the element but different mass numbers and number of neutrons what does the number next to isotopes signify the number indicates the isotope s mass number

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