

CONCEPTUAL Physics PRACTICE PAGE**Chapter 4 Newton's Second Law of Motion**
Mass and Weight

Learning physics is learning the connections among concepts in nature, and also learning to distinguish between closely-related concepts. Velocity and acceleration, previously treated, are often confused. Similarly in this chapter, we find that mass and weight are often confused. They aren't the same! Please review the distinction between mass and weight in your textbook. To reinforce your understanding of this distinction, circle the correct answers below.



Comparing the concepts of mass and weight, one is basic—fundamental—depending only on the internal makeup of an object and the number and kind of atoms that compose it. The concept that is fundamental is: [mass] [weight].

The concept that additionally depends on location in a gravitational field is: [mass] [weight].

[Mass] [Weight] is a measure of the amount of matter in an object and only depends on the number and kind of atoms that compose it.

It can correctly be said that [mass] [weight] is a measure of "laziness" of an object.

[Mass] [Weight] is related to the gravitational force acting on the object.

[Mass] [Weight] depends on an object's location, whereas [mass] [weight] does not.

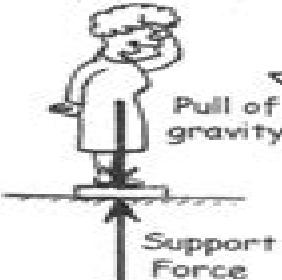
In other words, a stone would have the same [mass] [weight] whether it is on the surface of Earth or on the surface of the Moon. However, its [mass] [weight] depends on its location.

On the Moon's surface, where gravity is only about $1/6^{\text{th}}$ Earth gravity [mass] [weight] [both the mass and the weight] of the stone would be the same as on Earth.

While mass and weight are not the same, they are [directly proportional] [inversely proportional] to each other. In the same location, twice the mass has [twice] [half] the weight.

The Standard International (SI) unit of mass is the [kilogram] [newton], and the SI unit of force is the [kilogram] [newton].

In the United States, it is common to measure the mass of something by measuring its gravitational pull to Earth, its weight. The common unit of weight in the U.S. is the [pound] [kilogram] [newton].



When I step on a weighing scale, two forces act on it: a downward pull of gravity, and an upward support force. These equal and opposite forces effectively compress a spring inside the scale that is calibrated to show weight. When in equilibrium, my weight = mg .

Conceptual Physical Science Exploration Answers

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web video answer so now we ll talk about problems 58 from chapter four this problem whereas how do isotopes differ and how are they similar isotopes for the same element so isotopes differ in the number of neutrons

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web isotopes are atoms of the same element that differ in the amount of neutrons and atomic mass almost all elements on the periodic table have at least two different natural isotopes many elements

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web the isotopes of a given element always contain the same number of protons and therefore occupy the same place on the periodic table however because isotopes contain different numbers of neutrons each isotope has a unique atomic mass

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web apr 18 2015 isotopes are variants of a particular chemical element which differ in neutron number although all isotopes of a given element have the same number of protons in each atom answer link

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web jul 25 2014 isotopes differ in the number of neutrons in ions the number of electrons is different from the number of protons isotopes are atoms that have the same number of protons but different numbers of neutrons thus atoms of ^{61}C and of ^{63}C are isotopes of each other

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web isotopes are those atoms having same atomic number number of protons are same but different mass number number of neutrons differ ions are atoms which contain an overall charge where number of protons number of electrons

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web atoms of the same element i e same atomic number z that have different numbers of neutrons are called isotopes for example 99 of the carbon atoms on earth have 6 neutrons and 6 protons in their nuclei about 1 of the carbon atoms have 7

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web aug 19 2022 atoms with the same number of protons but different numbers of neutrons are called isotopes they share almost the same chemical properties but differ in mass and therefore in physical properties there are stable isotopes which do not emit radiation and there are unstable isotopes which do emit radiation

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web aug 9 2023 how do different isotopes of an element differ two isotopes differ in the number of neutrons in the nucleus for example carbon 14 has 8 neutrons differing from carbon 12 which has 6

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web atoms of the same element that differ in their numbers of neutrons are called isotopes many isotopes occur naturally usually one or two isotopes of an element are the most stable and common different isotopes of an element generally have the same physical and chemical properties because they have the same numbers of protons and electrons

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web jan 5 2016 explanation isotopes are atoms of the same element same atomic number ie same number of protons in the nucleus but different atomic masses different mass numbers due to different numbers of neutrons in the nucleus for example we may get an isotope of chlorine with 17 protons and 18 neutrons hence atomic mass 35

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web may 16 2018 explanation an atom is characterized as a particular element by the number of positively charged massive nuclear particles i e z z 1 we gots hydrogen z 2 we gots helium z 3 lithium z 47 we gots silver the nucleus can also contain neutrons massive nuclear particles of zero charge

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web atoms of the same element with different numbers of neutrons are called isotopes close isotope atoms of an element with the same number of protons and electrons but different numbers of neutrons

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web oct 31 2023 an isotope is one of two or more species of atoms of a chemical element with the same atomic number and

position in the periodic table and nearly identical chemical behavior but with different atomic masses and physical properties every chemical element has one or more isotopes

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web jul 18 2018 answer link due to the number of neutrons isotopes differ with respect to the number of neutrons their nuclei contain where neutrons are massive particles of zero charge a nucleus may contain variable numbers of neutrons and this gives rise to the phenomenon of isotopes

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web isotopes are distinct nuclear species or nuclides as technical term of the same chemical element they have the same atomic number number of protons in their nuclei and position in the periodic table and hence belong to the same chemical element but differ in nucleon numbers mass numbers due to different numbers of neutrons in

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web verified by topper correct option is d atoms having same atomic number but different atomic mass are known as isotopes the difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons

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web part i answer the questions based on the above reading what is an isotope isotopes are versions of the same element they have the same number of protons and electrons as the element but different mass numbers and number of neutrons what does the number next to isotopes signify the number indicates the isotope's mass number

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web nov 27 2016 explanation basically isotopes are nuclei that have the same number of protons and are thus the same element but have unequal numbers of neutrons answer link the nucleus differs as a result of the number of neutrons in the atom the actual structure doesn't differ all that much but its mass does differ thus its density is a little

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