

Post-lab questions:

1. Using the volumes of HSO_3^- solution added for each of the three titrations you conducted (A, B, and C), calculate the number of moles HSO_3^- required to reach the endpoint in each titration.
 2. Using the moles of HSO_3^- calculated in #1 above, calculate the number of moles of MnO_4^- titrated according to the balanced equations you submitted in the prelab exercises.
 3. Using the moles of MnO_4^- you calculated in #2 above, calculate the starting concentration of permanganate ion for each of the three titrations.
 4. Write the electron configurations for Mn, Mn^{2+} , Mn^{4+} , Mn^{6+} , and Mn^{7+} .
 5. Sodium bismuthate (NaBiO_3) oxidizes MnCl_2 to MnO_4^- and is consequently reduced to Bi(OH)_3 . How many grams of MnCl_2 will be oxidized by 7.5 g of NaBiO_3 ?
 6. Balance the following oxidation-reduction reactions:
 - a. $\text{Mn}^{2+} + \text{BiO}_3^- \longrightarrow \text{MnO}_4^- + \text{Bi}^{3+}$ (in acid)
 - b. $\text{NO}_3^- + \text{I}_2 \longrightarrow \text{NO}_2 + \text{IO}_3^-$ (in acid)
-
- c. $\text{ClO}_3^- + \text{S}^{2-} \longrightarrow \text{Cl}^- + \text{S}$ (in base)
 - d. $\text{Mo(OH)}_3 + \text{HO}_2^- \longrightarrow \text{MoO}_4^{2-}$ (in base)
7. In each of the titrations performed in this experiment, the end point was apparent due to a color change in the material being oxidized or reduced. Suppose that none of the materials involved in the redox reaction are colored. Suggest a way one could determine the endpoint in such a case. (HINT: Think about acid-base titrations.)

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Chemical Education Material Study



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Table of Contents Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers

1. Understanding the eBook Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - The Rise of Digital Reading Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Personalized Recommendations

- Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers User Reviews and Ratings
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers and Bestseller Lists
5. Accessing Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers Free and Paid eBooks
- Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers Public Domain eBooks
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers eBook Subscription Services
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers Budget-Friendly Options
6. Navigating Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers Compatibility with Devices
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Highlighting and Note-Taking Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Interactive Elements Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
8. Staying Engaged with Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers

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9. Balancing eBooks and Physical Books Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Setting Reading Goals Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Fact-Checking eBook Content of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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