

Note that the end vowel of the Greek prefix is occasionally omitted from the spelling for clarity in pronunciation.

Hydrates

Prefixes are also used for the naming of hydrates, inorganic salts in which water molecules are a part of the crystalline structure of the solid. The waters of hydration (also called waters of crystallization) are bound to the ions of the salt and, in most cases, can be removed with the application of heat (see *Experiment 5*). For example, barium chloride is often purchased as barium chloride anhydrate: BaCl_2 . The formula of iron(III) chloride hexahydrate is $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$. The formula for Epsom salt (Figure 5.1) is $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$, magnesium sulfate heptahydrate.

Binary Acids

A binary acid is an aqueous solution of a compound formed by hydrogen and a more electronegative nonmetal. To name the acid, the prefix *Hydro-* and suffix *-ic* are added to the root of the nonmetal name.

Formula	Name
$\text{HCl}(aq)$	Hydrochloric acid
$\text{HBr}(aq)$	Hydrobromic acid
$\text{H}_2\text{S}(aq)$	Hydro sulfide acid

Writing Formulas from Names

In writing formulas for compounds, it is mandatory that the sum of the oxidation numbers of the elements equals zero (see *Dry Lab 2A*, rule 8). The name of the compound must dictate how the formula is to be written.

Example D2B.1 Write the formula for calcium nitride.

Calcium is in group 2A and therefore has an oxidation number of +2: Ca^{2+} . The nitrogen atom is always -3 when combined with a metal: N^{3-} . Therefore, for the sum of the oxidation numbers to equal zero, there must be three Ca^{2+} (a total of +6) for every two N^{3-} (a total of -6); the formula is Ca_3N_2 .

DRY LAB PROCEDURE

Procedure Overview: Given the formula of the compound, the proper names for a large number of compounds are to be written. Given the name of the compound, the formulas for a large number of compounds also are to be written.

Your instructor will assign the exercises you are to complete. Answer them on a separate piece of paper. Be sure to indicate the date, your lab section, and your desk number on your *Report Sheet*. Use the rules that have been described.

1. Name the following binary salts of the representative elements.
 - a. Na_3P
 - b. CaC_2
 - c. Ca_3P_2
 - d. Ca_2P_3
 - e. Ca_3P_5
 - f. Ca_2P_7
 - g. Ca_3P_9
 - h. Ca_4P_{11}
 - i. Ca_5P_{14}
 - j. Ca_6P_{17}
 - k. Ca_7P_{20}
 - l. Ca_8P_{23}
 - m. Ca_9P_{26}
 - n. $\text{Ca}_{10}\text{P}_{29}$
 - o. $\text{Ca}_{11}\text{P}_{32}$
 - p. $\text{Ca}_{12}\text{P}_{35}$
 - q. $\text{Ca}_{13}\text{P}_{38}$
 - r. $\text{Ca}_{14}\text{P}_{41}$
 - s. $\text{Ca}_{15}\text{P}_{44}$
 - t. $\text{Ca}_{16}\text{P}_{47}$
 - u. $\text{Ca}_{17}\text{P}_{50}$
 - v. $\text{Ca}_{18}\text{P}_{53}$
 - w. $\text{Ca}_{19}\text{P}_{56}$
 - x. $\text{Ca}_{20}\text{P}_{59}$
 - y. $\text{Ca}_{21}\text{P}_{62}$
 - z. $\text{Ca}_{22}\text{P}_{65}$
2. Name the following salts according to the old -ic, -ous system and the Stock system.

a. CrS	b. Cr_2O_3	c. $\text{Cr}_2\text{O}_7 \cdot 6\text{H}_2\text{O}$	d. CrCl_3	e. Cr_2Cl_3	f. $\text{CrCl}_3 \cdot 6\text{H}_2\text{O}$	g. HgCl_2	h. Hg_2Cl_2	i. HgO	j. Fe_2O_3	k. Fe_2O_4	l. FeS	m. CoO	n. $\text{CoBr}_2 \cdot 6\text{H}_2\text{O}$	o. SeF_4	p. SeO_2	q. Cu_2O	r. $\text{Fe}(\text{OH})_3$
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Alvin Strickler,1923 [Basic Concepts of Chemistry](#) Leo J. Malone,Theodore Dolter,2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance

Exploring General Chemistry in the Laboratory Colleen F. Craig,Kim N. Gunnerson,2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science **Organic Chemistry I Laboratory**
Matt Davies,2015-11-11 *A Laboratory Manual of Organic Chemistry for Beginners* Arnold Frederick Holleman,1913

Journal of Chemical Education ,1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers **Integrated Physics and Chemistry Lab Manual Answer Key, 3rd Edition** Bright Thinker,2024-08 Key Individual Lab Manual Answer Key for Integrated Physics and Chemistry IPC 3rd edition **LAB MANUAL FOR CHEMISTRY: ATOMS FIRST** Gregg Dieckmann,John W. Sibert,2017-10-31 This laboratory manual presents a curriculum that is organized around an atoms first approach to general chemistry Our motivation for writing this manual is to 1 tap into the natural curiosity present in all of us and provide engaging experiments that students will find interesting 2 emphasize topics that students find particularly challenging in the general chemistry lecture course and 3 create a laboratory environment that encourages students on occasion to solve puzzles and not just follow recipes All too often students view general chemistry lab as a boring exercise in which an exact set of instructions is followed leading to an answer that in many cases results in a good grade regardless of how much learning has taken place To these students the successful lab is the one that takes the least amount of time Unfortunately a huge opportunity to get students truly turned on to science is missed To us the laboratory represents high stakes ground for

engagement and relatively low stakes for grading as the laboratory is typically a single credit course or minor component to the lecture grade. Thus while the rigor of the experiments in this manual can be tuned to meet the needs of the instructor our hope is that students will be encouraged to play safely with chemical concepts and laboratory techniques with grades simply being a natural consequence of their laboratory actions. To facilitate such a mindset this manual has been written to provide instructors with a weekly tool that can attract and keep student interest while providing important connections to the material covered in an atoms first lecture course. Our philosophy student curiosity leads to engagement which leads to discovery which leads to learning. The manual is for a freshman level general chemistry laboratory course and serves as an ideal supplement for any atoms first general chemistry textbook such as Chemistry Atoms First by Julia Burdge and Jason Overby. It is designed for students at all levels from those seeing chemistry for the first time to chemistry majors.

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