

Chem II Final Exam Study Guide

Lecture Notes

Chapter 14 - Solutions

Mixtures:

- Dissolve vs dissociate: dissolve means to break apart in a solution, dissociate means to break the molecule up into ions in a solution
- Entropy (S) = a drive to disorder
 - Mixtures are more disordered than pure substances.

Intermolecular forces:

- a. Definition: interaction between individual molecules
- b. To break IMFs = release energy (endothermic)
- c. To build IMFs = to absorb energy (endothermic)
- d. Characteristics of strong IMFs:
 - i. Higher boiling points, lower melting points, higher viscosity, lower MP
- e. 4 types of IMFs:
 - i. Dispersion: in everything. Dispersion is **stronger in larger molecules**
 - ii. Dipole-dipole: occurs in polar molecules, higher ΔH = higher d-d forces
 - iii. H bonds: hydrogen with N, O, F, occurs in polar molecules
- f. Polar dissolves polar, non polar dissolves non polar.
 - i. Less nonpolar shrubbery = more polar
- g. Increasing / decreasing volume in a solution = looser, weaker bonds
- h. Examples of solutions:
 - i. Ideal solution: breaks the same type of bonds that were there originally $\Delta V = 0$, $\Delta T = 0$, $\Delta H = 0$
 - ii. Miscible non-ideal: break weaker, form stronger, $\Delta V = (-)$ decrease, $\Delta T = (+)$ increase, $\Delta H = (-)$ exothermic
 - iii. Tends to be immiscible, non ideal: ex: oil + water , break stronger IMFs, form weaker, $\Delta V = (+)$ increase $\Delta T = (-)$ decrease, $\Delta H = (+)$ endothermic

Concentration units:

- a. m = molality = mol solu / Kg solv
- b. M = molarity mol solu / L soln
- c. % m/v = g solu / mL soln
- d. ppm/ ppb/ ppt = 1 g solu / 10^6 g soln , 10^3 for ppb, 10^{12} for ppt
- e. Mole fraction: $X_a = \frac{n_a}{n_a + n_b}$

Henry's Law:

- Solubility of a gas when you're trying to dissolve it into a liquid
- $S = K_H \times P_{gas}$ KH is given, pressure in ATM

Raoult's Law:

- **Use everything you see vapor pressure**
- Tendency for a mixture to evaporate

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