

## Answer Key for Chemical Equilibria

1. Equilibrium constant expressions are always written as showing a ratio of products over reactants. Here is a general format for a generic equilibrium:



The equilibrium constant,  $K_c$ , has the ratio expressed as molar concentrations:

$$K_c = \frac{[A]^a[B]^b}{[C]^c[D]^d}$$

The equilibrium constant,  $K_c$ , is called an analytical equilibrium constant and it is applied to an analytical equilibrium constant expression. The analytical equilibrium constant expression will only include those components in the equilibrium that are in the same phase. Solids are never included because one cannot measure the concentration of a solid.

The equilibrium constant,  $K_p$ , is applied to gas phase equilibria. The equilibrium constant expression is expressed as a ratio in terms of the partial pressures of the gaseous components. Applying this to the same equilibrium above, assuming A, B, C and D are all gases:

$$K_p = \frac{P_C^c \cdot P_D^d}{P_A^a \cdot P_B^b}$$

Use these concepts to write the equilibrium constant expressions for the equilibria in this question:



This is a homogeneous equilibrium, meaning all the components are in the same phase, in this case, the gas phase. Thus, we can write both a  $K_c$  expression and a  $K_p$  expression:

$$K_c = \frac{[N_2H_4]}{[N_2][H_2]^2} \quad \text{and} \quad K_p = \frac{P_{N_2H_4}}{P_{N_2} \cdot P_{H_2}^2}$$



This is also a homogeneous equilibrium, thus we can write both a  $K_c$  expression and  $K_p$  expression:

$$K_c = \frac{[PF_2][F_2]}{[PF_3]} \quad \text{and} \quad K_p = \frac{P_{PF_2} \cdot P_{F_2}}{P_{PF_3}}$$



This is a heterogeneous equilibrium. Solids are never included in the equilibrium constant expressions, so we will only include the gases in this equilibrium. Since gases are involved, we are able to write both a  $K_c$  and  $K_p$  equilibrium constant expression:

$$K_c = \frac{[Cl_2]^2}{[SiCl_4][H_2O]^2} \quad \text{and} \quad K_p = \frac{P_{Cl_2}^2}{P_{SiCl_4} \cdot P_{H_2O}^2}$$

# Chemical Equilibrium Expression Answer Key

**Nihon Kagakkai**

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## **Table of Contents Chemical Equilibrium Expression Answer Key**

1. Understanding the eBook Chemical Equilibrium Expression Answer Key
  - The Rise of Digital Reading Chemical Equilibrium Expression Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equilibrium Expression Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Equilibrium Expression Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Equilibrium Expression Answer Key
  - Personalized Recommendations
  - Chemical Equilibrium Expression Answer Key User Reviews and Ratings
  - Chemical Equilibrium Expression Answer Key and Bestseller Lists
5. Accessing Chemical Equilibrium Expression Answer Key Free and Paid eBooks
  - Chemical Equilibrium Expression Answer Key Public Domain eBooks
  - Chemical Equilibrium Expression Answer Key eBook Subscription Services

- Chemical Equilibrium Expression Answer Key Budget-Friendly Options
- 6. Navigating Chemical Equilibrium Expression Answer Key eBook Formats
  - ePUB, PDF, MOBI, and More
  - Chemical Equilibrium Expression Answer Key Compatibility with Devices
  - Chemical Equilibrium Expression Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Equilibrium Expression Answer Key
  - Highlighting and Note-Taking Chemical Equilibrium Expression Answer Key
  - Interactive Elements Chemical Equilibrium Expression Answer Key
- 8. Staying Engaged with Chemical Equilibrium Expression Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Equilibrium Expression Answer Key
- 9. Balancing eBooks and Physical Books Chemical Equilibrium Expression Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Equilibrium Expression Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Equilibrium Expression Answer Key
  - Setting Reading Goals Chemical Equilibrium Expression Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Equilibrium Expression Answer Key
  - Fact-Checking eBook Content of Chemical Equilibrium Expression Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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