

CHEMICAL EQUILIBRIUM

➤ $K_c = \frac{[B]}{[A]} = \text{mol L}^{-1}$

➤ $K_p = \frac{P_B}{P_A} = \text{Partial Pressure}$

➤ $K_x = \frac{X_B}{X_A} = \text{mole fraction}$

➤ $K_p = K_c(RT)^{\Delta n_g}$ where $\Delta n_g = n_p - n_r$

➤ Predicting the extent of reaction :

- $K_c > 10^3$ [Forward reaction is favoured.]
- $K_c < 10^{-3}$ [Reverse reaction is favoured.]
- $10^{-3} < K_c < 10^3$ [Both reactants and products are present in equilibrium]

➤ Free Energy Change (ΔG)

- If $\Delta G = 0$ then reversible reaction would be in equilibrium, $K_c = 1$
- If $\Delta G = (+)$ ve then equilibrium will be displaced in backward direction; $K_c < 1$
- If $\Delta G = (-)$ ve then equilibrium will shift in forward direction; $K_c > 1$

➤ (a) K_c unit $\rightarrow (\text{moles/lit})^{\Delta n_g}$

➤ ΔG or a reaction under any set of conditions is related to its value under standard conditions, i.e. ΔG° by the equation

$$\Delta G = \Delta G^\circ + 2.303 RT \log Q$$

➤ Under equilibrium conditions, for same number of moles of reactants and products

$$Q = K_p = K_c = K \text{ and } \Delta G = 0$$

\therefore

$$\Delta G^\circ = -2.303 RT \log K$$

➤ We have replaced K_p by K called thermodynamic equilibrium constant

Significance of ΔG°

Significance of ΔG° can be explained from the following points

(i) If $\Delta G^\circ < 0$, $\log K > 0$, $\Rightarrow K > 1$

Hence, reaction is spontaneous in forward direction.

(ii) Hence $\Delta G^\circ < 0$, $\log K > 0$, $\Rightarrow K < 1$

Hence, reaction is non-spontaneous or a reaction proceeds in the forward direction to such a small extent that a very small amount of the product is formed.



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