

Chemistry Benchmark II Study Guide

1. Write the formula for the following:
  - iron (II) oxide
  - calcium phosphate
  - ammonium phosphate
  - copper (III) bromide
  - nitrogen dioxide
  - titantium (IV) oxide
  - potassium bromide
  - magnesium hydroxide
  - xenon hexafluoride
2. What is a binary compound? What is a ternary compound?
3. Write formulas for:
  - hydroiodic acid
  - carbonic acid
  - hydrosulfuric acid
  - hydrochloric acid
  - hypochlorous acid
  - perchloric acid
  - phosphorous acid
  - chloric acid
  - sulfuric acid
  - chlorous acid
4. Name these:
  - NaCl
  - KI
  - CaBr
  - Rb<sub>2</sub>O
  - Sr<sub>2</sub>
  - CuCl
  - HgO
  - Fe<sub>2</sub>O<sub>3</sub>
  - MnO<sub>2</sub>
  - PbCl<sub>4</sub>
  - SnBr<sub>4</sub>
  - BF<sub>3</sub>
  - NO
  - N<sub>2</sub>O<sub>5</sub>
  - CCl<sub>4</sub>
  - P<sub>2</sub>O<sub>5</sub>
  - N<sub>2</sub>O<sub>3</sub>
5. Find the oxidation number assigned to:
  - S in H<sub>2</sub>SO<sub>4</sub>
  - P in PCl<sub>5</sub>
  - S in SO<sub>3</sub>
6. Name these:
  - NO<sub>2</sub>
  - N<sub>2</sub>O
  - SO<sub>3</sub>
7. How many valence electrons do elements have in the following groups?
  - alkali metals
  - alkaline earth metals
  - halogens
  - noble gases
8. Write the complete electron configuration, noble gas configuration, Lewis dot structure, and orbital notation for:
  - Sr
  - Br
  - Na
  - K
  - N
  - Cl
9. In each group, identify the atom with the largest atomic radius.
  - Sn, Xe, Rb, Sr
  - IRn, He, Xe, Kr
  - Pb, Ba, Cs, At
10. In each group, identify the atom with the highest ionization energy.
  - Cs, K, Li
  - Ba, Sr, Ca
  - I, Br, Cl
  - Mg, Si, S
11. Define the following: ionic bond, covalent bond, polar covalent bond, nonpolar covalent bond.
12. Define electronegativity and describe the trend across a period and down a group. What type of element has the highest electronegativity? What type has the lowest electronegativity?
13. Identify the type of bond in each:
  - K-N
  - Cs-O
  - C-N
  - O-F
  - Na-Cl
  - Cl-F
  - Cl-O
  - Ca-Cl
14. Who first arranged the periodic table in order of increasing atomic mass?
15. Draw Lewis structures for the following and predict the molecular shape.
  - H<sub>2</sub>S
  - PH<sub>3</sub>
  - O<sub>2</sub>
  - sulfate ion
  - SO<sub>3</sub>
  - BF<sub>3</sub>
16. Calculate the wavelength and energy of 96.6MHz radio waves.
17. What is the frequency and energy of electromagnetic radiation having a wavelength of  $3.33 \times 10^{-9}$ m?
18. Predict the ions that the following elements might form:
  - F
  - K
  - I
  - Ge
  - Fe (trick question)
  - P
  - Ag

# Chemistry Benchmark Study Guide Answer Key

**United States. Environmental  
Protection Agency**

## Chemistry Benchmark Study Guide Answer Key:

*Study Guide for Bailey and Scott's Diagnostic Microbiology - E-Book* Betty A. Forbes, Daniel F. Sahm, Alice S. Weissfeld, 2016-06-30 Corresponding to chapters in Bailey Scott's Diagnostic Microbiology 12th Edition this new guide reviews important topics and helps students master key material. It includes chapter objectives a summary of key points review questions and case studies. Material is presented in an engaging format that challenges students to apply their knowledge to real life scenarios. Type Source Promotion Chapter Objectives open each chapter providing a measurable outcome to achieve by completing the material. A summary of Key Points from the main text helps students clearly identify key concepts covered in each chapter. Review Questions in each chapter test students on important knowledge in addition to key terms and abbreviations. Case studies in each chapter offer challenging questions for further analysis and challenge students to apply their knowledge to the real world.

**Scientific and Technical Books and Serials in Print**, 1984 **Findex**, 1985 **Books in Print Supplement**, 1985 **EPA Publications Bibliography** United States. Environmental Protection Agency, 1996 **Resources in Education**, 1984 **Scientific and Technical Aerospace Reports**, 1994 **EPA Publications Bibliography Quarterly Abstract Bulletin** United States. Environmental Protection Agency, 1996-07 **Bowker's Complete Video Directory 2001**, 2001 **EPA Publications Bibliography**, 1996 **Forthcoming Books** Rose Arny, 2001-06 **Teaching Science in Elementary and Middle School Classrooms** Joseph S. Krajcik, Charlene Lochbihler Czerniak, Carl F. Berger, 2003 This text provides an overview of current science teaching practices for the elementary and middle grades. The authors top scholars in the field of science education believe that all children should develop an in depth and meaningful understanding of scientific concepts and processes. To achieve this the text utilizes the Project Based Approach. Project based science stresses that science teaching should emphasize the active engagement of students in science rather than teachers telling students information. Each chapter has several Portfolio Activity boxes that provide active learning experiences or reflections for the student. Like the first edition the text includes numerous strategies in each chapter that help both new and experienced teachers understand how to teach science in an active and engaging manner. The text also shows teachers how to implement the National Science Education Standards NSES and constructivist strategies. A NSES marginal feature keys content to the standards. Moreover this textbook helps teachers learn how to implement all of today's major reforms not just read about them.

**The Publishers' Trade List Annual**, 1986 **Technical Book Review Index**, 1976 **Resources in Education**, 1997 **Bibliography of Agriculture**, 1992-11 **Subject Guide to Books in Print**, 1993 **Previews**, 1979 **Plating and Surface Finishing**, 1998 **El-Hi Textbooks and Serials in Print**, 1995 R R Bowker Publishing, 1995-02

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