



CHEMISTRY CHEAT SHEET



EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol K_c .

What is Equilibrium?		Equilibrium Constant	Graphical Representation of Equilibrium
<p>What is Equilibrium?</p> <p>Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol K_c.</p> <p>Equilibrium Constant</p> <p>$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$</p>	<p>Equilibrium Constant</p> <p>$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$</p>	<p>Graphical Representation of Equilibrium</p>	
<p>Equilibrium Constant, K_c</p> <p>Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol K_c.</p> <p>$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$</p>	<p>Equilibrium Constant, K_c</p> <p>$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$</p>	<p>Equilibrium Constant, K_c</p> <p>$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$</p>	
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Le Chatelier's Principle: the effect of changes in conditions on equilibrium

Change	Effect on Equilibrium	Effect on K_c
Concentration	<ul style="list-style-type: none"> Increasing the concentration of a reactant or product will shift the equilibrium position to the opposite side. 	<p>K_c is not affected by changes in concentration.</p>
Pressure	<ul style="list-style-type: none"> Increasing the pressure will shift the equilibrium position to the side with the fewest moles of gas. 	<p>K_c is not affected by changes in pressure.</p>
Temperature	<ul style="list-style-type: none"> Increasing the temperature will shift the equilibrium position to the endothermic side. 	<p>K_c is affected by changes in temperature.</p>

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Susan McMurry



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Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

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Student Study Guide to accompany Foundations of College Chemistry, 11th Edition Morris Hein, Susan Arena, 2003-09-03 Work more effectively and gauge your progress along the way This Student Study Guide that accompanies Hein's Foundations of College Chemistry 11th Edition contains a review and self evaluation section challenge problems and answers and solutions for all exercises in the study guide Foundations of College Chemistry is the book that defined the prep intro chemistry market 35 years ago And it's been a market leader ever since Hein Arena is known for its accuracy clear no nonsense approach and direct writing style Strong problem solving and carefully constructed problem sets make this book a standout among its many imitators

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Study Guide & Solutions Manual to Accompany Organic Chemistry, Third Edition G. Marc Loudon, Joseph G. Stowell, 1995

Resources in Education, 1984

Study Guide and Solutions Manual for Fundamentals of General, Organic, and Biological Chemistry Susan McMurry, 1992

Study Guide for Hein and Arena's Foundations of College Chemistry Peter Scott, Rachael Enriquez Porter, 1999-04-18

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variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that s real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You ll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

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Fundamental Concepts of Applied Chemistry Jayashree Ghosh,2010-12

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