

Chemistry Spring Final Study Guide

Unit 10: Acids & Bases

Use the following information and complete the structure of Acids & Bases.

Acids	Bases
<p>Unit 10: Acids & Bases</p> <p>Draw a line on the graph and indicate its dependence for each property: Temperature, Volume, and Pressure below.</p> <p>Label the y-axis.</p> <p>Label the x-axis.</p> <p>1. A substance of the particles.</p> <p>2. A substance of the particles.</p> <p>3. A substance of the particles.</p> <p>4. The gas is in.</p> <p>5. The volume.</p>	<p>Unit 10: Acids & Bases</p> <p>Draw a line on the graph and indicate its dependence for each property: Temperature, Volume, and Pressure below.</p> <p>Label the y-axis.</p> <p>Label the x-axis.</p> <p>1. A substance of the particles.</p> <p>2. A substance of the particles.</p> <p>3. A substance of the particles.</p> <p>4. The gas is in.</p> <p>5. The volume.</p>

Unit 11: Chemical Reactions

Write the balanced chemical equation for each reaction.

1. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

2. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$

3. $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

4. $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$

5. $\text{Ca} + \text{O}_2 \rightarrow \text{CaO}$

6. $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$

7. $\text{K} + \text{O}_2 \rightarrow \text{K}_2\text{O}$

8. $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$

9. $\text{Li} + \text{O}_2 \rightarrow \text{Li}_2\text{O}$

10. $\text{Ba} + \text{O}_2 \rightarrow \text{BaO}$

11. $\text{Sr} + \text{O}_2 \rightarrow \text{SrO}$

12. $\text{Ca} + \text{O}_2 \rightarrow \text{CaO}$

13. $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$

14. $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$

15. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$

16. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

17. $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

18. $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

19. $\text{C}_5\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

20. $\text{C}_6\text{H}_{14} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

Unit 12: Chemical Reactions

Write the balanced chemical equation for each reaction.

1. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

2. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$

3. $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

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11. $\text{Sr} + \text{O}_2 \rightarrow \text{SrO}$

12. $\text{Ca} + \text{O}_2 \rightarrow \text{CaO}$

13. $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$

14. $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$

15. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$

16. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

17. $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

18. $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

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Unit 13: Chemical Reactions

Write the balanced chemical equation for each reaction.

1. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

2. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$

3. $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

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R Bogdan



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