

ACS INORGANIC CHEMISTRY STUDY GUIDE

1. Period 2: O₂, F₂ etc, why not Ne₂? difference is MO diagram's relative energies of the bonding orbitals piu and pig
2. H-X: H is something that bonds through 2p and 2s orbitals, relative energies of the atomic orbitals are taken into account because there are two separate atoms, H 1s is usually higher in energy, 4 atomic orbitals of X will mix with 1s of H to give 5 molecular orbital. Anti bonding is mostly H in character while the other 4 are mostly X
3. X-Y: relative energies are taken into account, more electronegative=lower energy of orbitals
4. Relative energies of MOs: 4sigma>2pi>3sigma>1pi>2sigma>1sigma
5. Polyatomic MO: must use approximations of MO theory. Linear combinations of Atomic orbitals create a cumulative approximation of the atomic orbitals of its constituent atoms, compare orbitals to determine what will mix
6. LCAO notation: a,b=nondegenerate, e=doubly degenerate, t=triple degenerate
7. LCAO energies: lowest=a (nondegenerate) intermediate=e highest=t
8. LCAO: linear combinations of atomic orbitals

9. Bond order: a method to assessing overall bond strength between two atoms in a molecule, higher bond order=stronger bond, assesses net number of bonds between two atoms
10. Calculating bond order: (bonding-antibonding)/2
11. How to assign bonding character-bonding: a bonding orbital will lie lower in energy than its substituent atomic orbitals
12. How to assign bonding character-nonbonding: a nonbonding orbital will be equal in energy to its corresponding atomic orbitals
13. How to assign bonding character-antibonding: an antibonding orbital will lie higher in energy than its substituent atomic orbitals
14. HOMO interactions: highest occupied molecular orbital, highest energy electrons reside
15. LUMO interactions: lowest unoccupied molecular orbital, it is the energy level directly above the homo
16. Importance of HOMO/LUMO: critical to electron structure, when excited an electron is promoted from HOMO to LUMO, bonding and anti bonding characteristics determine stability, the frontier orbitals are the point of interaction and basis of understanding for molecular

College Inorganic Chemistry Study Guide Problems

Marco Cascella

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The United States Catalog Mary Burnham, Carol Hurd, 1928 *Organic Chemistry Study Guide with Solutions Manual* Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer [The United States Catalog](#) ,1906 **The American Catalog** ,1881
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Within the pages of "**College Inorganic Chemistry Study Guide Problems**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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