

- (d) Write the chemical equation of the reaction which takes place on reacting XNO_3 solution and common salt solution giving the physical states of all the reactants and products.
- (e) What type of chemical reaction is illustrated by the above equation?
64. Two metals X and Y form the salts XSO_4 and Y_2SO_4 respectively. The solution of salt XSO_4 is blue in colour whereas that of Y_2SO_4 is colourless. When barium chloride solution is added to XSO_4 solution, then a white precipitate Z is formed along with a salt which turns the solution green. And when barium chloride solution is added to Y_2SO_4 solution, then the same white precipitate Z is formed along with colourless common salt solution.
- (a) What could the metals X and Y be?
- (b) Write the name and formula of salt XSO_4 .
- (c) Write the name and formula of salt Y_2SO_4 .
- (d) What is the name and formula of white precipitate Z?
- (e) Write the name and formula of the salt which turns the solution green in the first case.
65. A red-brown metal X forms a salt XSO_4 . When hydrogen sulphide gas is passed through an aqueous solution of XSO_4 , then a black precipitate of XS is formed along with sulphuric acid solution.
- (a) What could the salt XSO_4 be?
- (b) What is the colour of salt XSO_4 ?
- (c) Name the black precipitate XS .
- (d) By using the formula of the salt obtained in (a) above, write an equation of the reaction which takes place when hydrogen sulphide gas is passed through its aqueous solution.
- (e) What type of chemical reaction takes place in this case?
66. When a strip of red-brown metal X is placed in a colourless salt solution YNO_3 , then metal Y is set free and a blue coloured salt solution $\text{X(NO}_3)_2$ is formed. The liberated metal Y forms a shining white deposit on the strip of metal X.
- (a) What do you think metal X is?
- (b) Name the salt YNO_3 .
- (c) What could be metal Y?
- (d) Name the salt $\text{X(NO}_3)_2$.
- (e) What type of reaction takes place between metal X and salt solution YNO_3 ?
67. A metal salt MX when exposed to light splits up to form metal M and a gas X_2 . Metal M is used in making ornaments whereas gas X_2 is used in making bleaching powder. The salt MX is itself used in black and white photography.
- (a) What do you think metal M is?
- (b) What could be gas X_2 ?
- (c) Name the metal salt MX .
- (d) Name any two salt solutions which on mixing together can produce a precipitate of salt MX .
- (e) What type of chemical reaction takes place when salt MX is exposed to light? Write the equation of the reaction.

ANSWERS

1. Decomposition reaction 4. $\text{Zn} + 2\text{AgNO}_3 \longrightarrow \text{Zn(NO}_3)_2 + 2\text{Ag}$ 5. Rancidity 6. Anti-oxidants
11. Decomposition reactions (carried out by electricity) 14. $2\text{AgNO}_3(\text{aq}) + \text{Cu}(\text{s}) \longrightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2\text{Ag}(\text{s})$ 15. (i) Decomposition (ii) Combination (iii) Decomposition (iv) Decomposition (v) Decomposition
- (c) Combination 16. (a) Combination (b) Decomposition (c) Decomposition (d) Decomposition
- (e) Combination 17. (i) Displacement reaction (ii) Combination reaction (iii) Decomposition reaction
- (iv) Double displacement reaction (v) Displacement reaction 18. (a) See page 29 (b) See page 29
19. (a) Displacement reaction (b) Combination reaction 20. (a) Combination reaction
- (b) Displacement reaction (c) Displacement reaction (d) Decomposition reaction (e) Double displacement reaction
21. (a) H_2O_2 (b) PbS 22. H_2S 23. Substance oxidised : H_2S ; Substance reduced : SO_2
24. (a) oxidation ; reduction (b) reduction ; oxidation (c) rancidity 25. (i) C (ii) ZnO 26. (b) Magnesium is oxidised ; Addition of oxygen to magnesium takes place (c) (i) HCl (ii) MnO_2 (iii) MnO_2 (iv) HCl
31. $\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}$ (i) Zinc oxide (ii) Carbon 34. (b) $2\text{Cu} + \text{O}_2 \longrightarrow 2\text{CuO}$ (i) Cu (ii) O_2
35. (b) $2\text{Mg} + \text{O}_2 \longrightarrow 2\text{MgO}$ (i) Mg (ii) O_2 37. (b) $\text{Cu}(\text{s}) + 2\text{AgNO}_3(\text{aq}) \longrightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2\text{Ag}(\text{s})$ 39. (a) H_2 (b) CuO (c) CuO (d) H_2 43. (b) Rusting (c) Oxidation 44. (b) Oxidation

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