

(a) Write the chemical equation of the reaction which takes place on reacting XNO_3 solution and common salt solution giving the physical states of all the reactants and products.
 (b) What type of chemical reaction is illustrated by the above equation?

64. Two metals X and Y form the salts XSO_4 and Y_2SO_4 , respectively. The solution of salt XSO_4 is blue in colour whereas that of Y_2SO_4 is colourless. When barium chloride solution is added to XSO_4 solution, then a white precipitate Z is formed alongwith a salt which turns the solution green. And when barium chloride solution is added to Y_2SO_4 solution, then the same white precipitate Z is formed alongwith colourless common salt solution.
 (a) What could the metals X and Y be?
 (b) Write the name and formula of salt XSO_4 .
 (c) Write the name and formula of salt Y_2SO_4 .
 (d) What is the name and formula of white precipitate Z?
 (e) Write the name and formula of the salt which turns the solution green in the first case.

65. A red-brown metal X forms a salt XSO_4 . When hydrogen sulphide gas is passed through an aqueous solution of XNO_3 , then a black precipitate of X_2S is formed alongwith sulphuric acid solution.
 (a) What could the salt XNO_3 be?
 (b) What is the colour of salt XNO_3 ?
 (c) Name the black precipitate X_2S .
 (d) By using the formula of the salt obtained in (a) above, write an equation of the reaction which takes place when hydrogen sulphide gas is passed through the aqueous solution.
 (e) What type of chemical reaction takes place in this case?

66. When a strip of red-brown metal X is placed in a colourless salt solution YNO_3 , then metal Y is set free and a blue coloured salt solution X_2NO_3 is formed. The liberated metal Y forms a shining white deposit on the strip of metal X.
 (a) What do you think metal X is?
 (b) Name the salt YNO_3 .
 (c) What could be metal Y?
 (d) Name the salt X_2NO_3 .
 (e) What type of reaction takes place between metal X and salt solution YNO_3 ?

67. A metal salt MX when exposed to light splits up to form metal M and a gas X_2 . Metal M is used in making ornaments whereas gas X_2 is used in making bleaching powder. The salt MX is itself used in black and white photography.
 (a) What do you think metal M is?
 (b) What could be gas X_2 ?
 (c) Name the metal salt MX.
 (d) Name any two salt solutions which on mixing together can produce a precipitate of salt MX.
 (e) What type of chemical reaction takes place when salt MX is exposed to light? Write the equation of the reaction.

ANSWERS

1. Decomposition reaction 4. $\text{Zn} + 2\text{AgNO}_3 \longrightarrow \text{Zn}(\text{NO}_3)_2 + 2\text{Ag}$ 8. Rancidity 6. Anti-oxidants
 11. Decomposition reactions carried out by electricity 14. $2\text{AgNO}_3 \text{ (aq)} + \text{Cu} \text{ (s)} \longrightarrow \text{Cu}(\text{NO}_3)_2 \text{ (aq)} + 2\text{Ag} \text{ (s)}$ 15. (i) Decomposition (ii) Combination 16. (i) Decomposition (ii) Decomposition
 (iii) Decomposition 17. (i) Displacement reaction (ii) Combination reaction 18. (a) See page 29 (b) See page 29 19. (a) Displacement reaction (b) Combination reaction 20. (a) Combination reaction
 (b) Displacement reaction (c) Displacement reaction (d) Decomposition reaction (e) Double displacement reaction 21. (a) H_2O_2 (b) PbS 22. H_2S 23. Substance oxidised : H_2S ; Substance reduced : SO_2
 24. (a) oxidation ; reduction (b) reduction ; oxidation (c) rancidity 25. (i) C (ii) ZnO 26. (a) Magnesium is oxidised ; Addition of oxygen to magnesium takes place (i) (i) HCl (ii) MnO_2 (iii) MnO_3 (iv) HCl
 27. $\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}$ (i) Zinc oxide (ii) Carbon 28. (i) $2\text{Cu} + \text{O}_2 \longrightarrow 2\text{CuO}$ (ii) O_2
 29. (i) $2\text{Mg} + \text{O}_2 \longrightarrow 2\text{MgO}$ (ii) Mg (iii) O_2 30. (i) Cu (ii) O_2 31. (i) Cu (ii) $2\text{AgNO}_3 \text{ (aq)} \longrightarrow \text{Cu}(\text{NO}_3)_2 \text{ (aq)} + 2\text{Ag} \text{ (s)}$ 32. (a) H_2 (b) CuO (c) H_2 33. (b) Rusting (c) Oxidation 34. (b) Oxidation

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