

# Unit 1 Chemistry Study Guide

## History of the Atom:

### Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

### Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

### Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

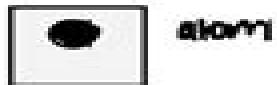
### Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

### John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
  - o **Daltons Atomic Theory:**
    - o All matter is composed of extremely small particles called atoms.
    - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
    - o Atoms cannot be subdivided, created, or destroyed.
    - o Atoms to different elements can combine in simple whole number ratios to form chemical compounds.
    - o In chemical reactions, atoms are combined, separated or rearranged.
  - o **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

### Billiard Ball Model:



### J.J. Thompson (1906):

- **Cathode Ray Experiment –**
  - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

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