

Chem 105: Exam 3 Study Guide

October 19th, 2020 Lecture

Basics of Quantum Mechanics:

Quantum Mechanics → focuses on electrons or photons.

Light

- is electromagnetic radiation
- is wave-like and particle-like
- is not continuous with infinite values, it can only have specific values (shown through visible spectrum)
- must overcome binding energy of an electron to eject it from a metal
- only has very specific values

Frequency threshold has to be met to achieve the photoelectric effect.

Electromagnetic Spectrum:

Visible light

- Violet → lowest visible wavelength
- Red → highest visible wavelength
- Lowest energy → highest energy

Each element gives off electrons at specific light energies

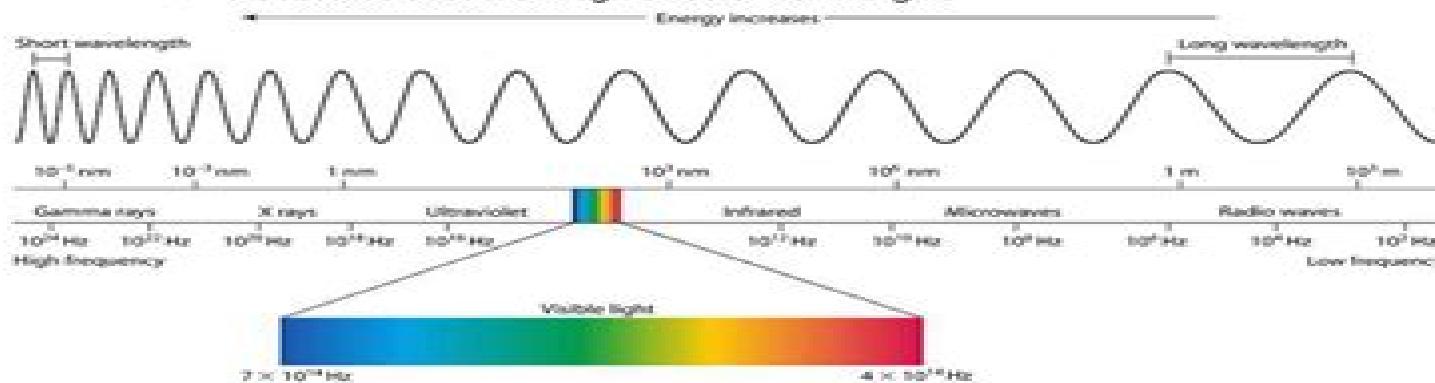
- can differentiate elements based on this
- light is not continuous because of this

The further out the electron is the higher energy it gives off

- High end → lower one : emit light
- Low level → higher one : absorbs light

Only a few light colors show up in visible light when reading colors for elements

- ROYGBIV
- the rest are shown through all other wavelengths



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