

Chapter 10: INTERMOLECULAR FORCES - Study Guide

CHEM 1412 Chemistry OpenStax

Assessment

1. Explain the difference between *intermolecular forces* and *intramolecular forces*.
2. While dipole forces are present in all substances, they are the most significant type of intermolecular force only in substances that are normally *nonpolar*.
Explain briefly what causes an instantaneous dipole to be formed.

3. "What is *polarity*?"

4. In which of the following molecules would dispersion forces be the primary type of intermolecular force?
CO Na+ NH3 CO2 KCl H2

5. Dipole-dipole interactions are forces of attraction between molecules which are permanently

6. Explain briefly what is meant by the term *dipole*, and why would they be attracted to one another.

7. Which of the following molecules would be expected to have significant dipole-dipole interactions in the liquid state?
CO Na+ NH3 CO2 KCl H2

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L Darling-Hammond

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