

Chemistry 162 Study Guide (first exam)

Useful equations: $S = k_B(P)$, $[A]_t = -kt + [A]_0$, $P_{\text{total}} = X_A(P_A) + X_B(P_B)$, $1/[A]_t = kt + 1/[A]_0$, $\Delta T = K_{\text{eb}}(m)$, $q = m(C_p)\Delta T$, $\ln[A]_t = -kt + \ln[A]_0$, $P_{\text{atmos}} = MRT$, $R = 0.08206 \text{ atm}\cdot\text{L/mol}\cdot\text{K}$, $T(\text{K}) = T(^{\circ}\text{C}) + 273.15$

1.) Which of the following is a liquid at room temperature (only one is a liquid)?

- a) CH_3OH b) CH_3CH_3 c) $\text{CH}_3\text{CH}_2\text{F}$ d) CH_3F e) CH_3OCH_3 f) CH_4

Intermolecular attractions hydrogen bonding is the strongest force

2.) Viscosity is the measure of the resistance of a liquid to flow, true or false? (Circle one)

No equations or anything just need to know what it is...

3.) What compound will **not** exhibit hydrogen bonding in the liquid state?

- a) $\text{CH}_3-\text{CH}_2-\text{Br}$ b) $\text{CH}_3-\text{CH}_2-\text{NH}_2$ c) $\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{OH}$ d) $\text{CH}_3-\text{NH}-\text{CH}_3$ e) $\text{NH}_2-\text{O}-\text{H}$

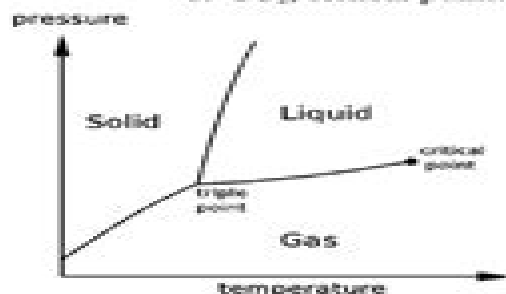
A is the only molecule which doesn't have X-H bond, X = N, O, or F

4.) Which compound cannot be liquefied by compression at temperature of 25.0°C ?

- a. CH_3Cl , critical point: 144.0°C , 66.0 atm
b. C_2H_2 , critical point: 35.5°C , 61.6 atm
c. C_2H_6 , critical point: 32.2°C , 48.2 atm

d. SiF_4 , critical point: -14.1°C , 36.7 atm. 25°C is beyond the critical point thus only a supercritical fluid is possible... see a phase diagram

- e. SO_2 , critical point: 158.0°C , 78.0 atm



5.) The solubility of O_2 in water is approximately 0.00380 g L^{-1} when the temperature is 25.0°C and the partial pressure of gaseous oxygen is 760 torr. What will the solubility of oxygen be if the oxygen pressure is readjusted to 1000 torr? Henry's law ($C = kP$) but the substance is the same so constant drops out.

- a. 0.00289 g L^{-1}

b. 0.00500 g L^{-1} Pg 490 $C_1/P_1 = C_2/P_2$ $C_2 = 0.0038(1000)/760 = 0.0050 \text{ g L}^{-1}$

- c. 1.49 g L^{-1}

- d. $2.89 \times 10^3 \text{ g L}^{-1}$

- e. $3.46 \times 10^3 \text{ g L}^{-1}$

6.) A solution is made by mixing 138.2 grams of ethanol, $\text{C}_2\text{H}_5\text{O}$, ($46.069 \text{ g mol}^{-1}$); 103.6 grams of water ($18.015 \text{ g mol}^{-1}$), and 80.11 grams of methanol, CH_3O ($32.042 \text{ g mol}^{-1}$). What is the mole fraction of methanol in the mixture? Calc mol of each 3 mol EtOH, 5.75 water, 2.5 MeOH.

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Mary H. Bailey



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Study Guide and Solutions Manual Neil E. Schore, 2002-08-02

Organic Chemistry Study Guide and Solutions

Manual Marc Loudon, Jim Parise, Lisa Bonner, 2024-08-09 The Study Guide and Solutions Manual offers several study tools for students Study Guide Links are brief supplements to text material and provide hints for study shortcuts approaches to problem solving or additional explanations of topics that typically cause difficulty Further Explorations are short discussions that provide additional depth on content in the text Reaction Reviews present the reactions in the chapter in detailed outline form This summary consolidates the reactions of the chapter in one place and presents not only the reactions themselves but also the essential features of their mechanisms and stereochemistry for cases in which these issues are discussed in the text Solutions are provided for all of the over 1600 problems in the text Many include worked out examples in which not only the answer is given but also the detailed approach to the solution is discussed

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Manual to Accompany The Essence of Organic Chemistry Jane M. Cram, Donald J. Cram, 1978

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Selected Solutions Manual for Basic Chemistry Karen C. Timberlake, 2016-03-29 The Study Guide and Selected Solutions Manual assists students with the text material It contains learning objectives chapter outlines additional problems with self tests and answers and answers to the odd numbered problems in the text

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Organic Chemistry, Third Edition G. Marc Loudon, Joseph G. Stowell, 1995

Study Guide and Solutions Manual for

McMurry's Organic Chemistry, Fifth Edition Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

Chemical Principles Student's Study Guide & Solutions

Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

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