

# AP Chemistry Final Study Guide

## Unit 1: Atomic Structure and Properties (7%-9%)

- At STP, one mole of gas always occupies 22.4 L
- **Mass Spectroscopy:**
  - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
  - The most abundant isotope will have the greatest effect
- **Coulomb's Law:**
  - The greater the charge of the nucleus, the more energy is required to ionize an electron
  - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- **Electromagnetic Radiation:**
  - When atoms absorb energy, electrons jump to higher energy levels
  - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- **Kinetic Energy of Ionized Electrons:**
  - The faster an ejected electron is going, the more kinetic energy it has
  - Electrons originally further away from the nucleus require less energy to eject, so it will have more kinetic energy
- **Electron Configuration:**
  - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
    - Nickel:  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$  (notice the d-block)
    - Nickel (with shorthand):  $[\text{Ar}]4s^2 3d^8$
- **Law of Definite Proportions:**
  - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- **Law of Multiple Proportions:**
  - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- **Aufbau Principle:**
  - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

# Chem Study Guide For Spring Benchmark

**ML Yell**

## **Chem Study Guide For Spring Benchmark:**

*Water quality in the Great and Little Miami River Basins, Ohio and Indiana, 1999-2001 , Energy Research*

**Abstracts** ,1993    **Chemists' Guide to Effective Teaching** Norbert J. Pienta,Melanie M. Cooper,Thomas J.

Greenbowe,2005 For courses in Methods of Teaching Chemistry Useful for new professors chemical educators or students learning to teach chemistry Intended for anyone who teaches chemistry or is learning to teach it this book examines applications of learning theories presenting actual techniques and practices that respected professors have used to implement and achieve their goals Each chapter is written by a chemist who has expertise in the area and who has experience in applying those ideas in their classrooms This book is a part of the Prentice Hall Series in Educational Innovation for Chemistry    *Innovative Techniques for Large-group Instruction* National Science Teachers Association,2002 Size does matter When you're faced with a class of 50 150 or even 250 college students it's tough to head off boredom much less promote higher order thinking and inquiry skills But it's not impossible thanks to the professor tested techniques in this collection of 14 articles from the Journal of College Science Teaching The book starts by examining what research shows about the effectiveness of popular teaching styles Surprise Lectures don't stimulate active learning From there the authors offer proven alternatives that range from small scale innovations to completely revamped teaching methods Suggested strategies include using quizzes in place of midterms and finals student forums interactive lectures collaborative groups group facilitators and e-mail and computer technology    [The Publishers' Trade List Annual](#) ,1985    **Resources in**

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