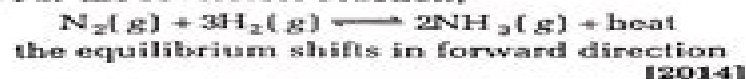


# 7

## Chemical Equilibrium

1. For the reversible reaction,



- (a) by increasing the concentration of  $\text{NH}_3(\text{g})$   
 (b) by decreasing the pressure  
 (c) by decreasing the concentrations of  $\text{N}_2(\text{g})$  and  $\text{H}_2(\text{g})$   
 (d) by increasing pressure and decreasing temperature

2. Using the Gibbs energy change,  $\Delta G^\circ = +63.3 \text{ kJ}$  for the following reaction,



the  $K_{\text{sp}}$  of  $\text{Ag}_2\text{CO}_3(\text{s})$  in water at  $25^\circ\text{C}$  is  
 ( $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$ ) [2014]

- (a)  $3.2 \times 10^{-26}$   
 (b)  $8.0 \times 10^{-12}$   
 (c)  $2.9 \times 10^{-3}$   
 (d)  $7.9 \times 10^{-2}$

3.  $\text{KMnO}_4$  can be prepared from  $\text{K}_2\text{MnO}_4$  as per reaction,



The reaction can go to completion by removing  $\text{OH}^-$  ions by adding [2013]

- (a)  $\text{HCl}$   
 (b)  $\text{KOH}$   
 (c)  $\text{CO}_2$   
 (d)  $\text{SO}_2$

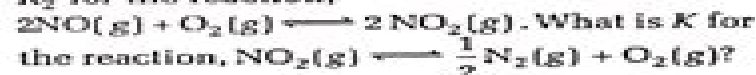
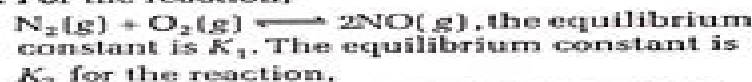
4. The value of  $\Delta H$  for the reaction,



is less than zero. Formation of  $\text{XY}_4(\text{g})$  will be favoured at [2011]

- (a) low pressure and low temperature  
 (b) high temperature and low pressure  
 (c) high pressure and low temperature  
 (d) high temperature and high pressure

5. For the reaction,



- (a)  $1/(4K_1K_2)$  (b)  $\{1/K_1K_2\}^{1/2}$  [2011]  
 (c)  $1/(K_1K_2)$  (d)  $1/4K_1K_2$

6. In which of the following equilibrium  $K_c$  and  $K_p$  are not equal? [2010]

- (a)  $2\text{NO}(\text{g}) \rightleftharpoons \text{N}_2(\text{g}) + \text{O}_2(\text{g})$   
 (b)  $\text{SO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons \text{SO}_3(\text{g}) + \text{NO}(\text{g})$   
 (c)  $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$   
 (d)  $2\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{CO}_2(\text{g})$

7. The dissociation constants for acetic acid and  $\text{HCN}$  at  $25^\circ\text{C}$  are  $1.5 \times 10^{-5}$  and  $4.5 \times 10^{-10}$ , respectively. The equilibrium constant for the equilibrium,



- (a)  $3.0 \times 10^5$  (b)  $3.0 \times 10^{-5}$   
 (c)  $3.0 \times 10^{-4}$  (d)  $3.0 \times 10^4$

# Chemical Equilibrium Chemistry Study Guide Answers

**Steven S. Zumdahl, Paul B. Kelter**



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