

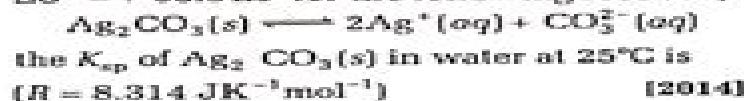
## 7

## Chemical Equilibrium

1. For the reversible reaction,



- (a) by increasing the concentration of  $NH_3(g)$
- (b) by decreasing the pressure
- (c) by decreasing the concentrations of  $N_2(g)$  and  $H_2(g)$
- (d) by increasing pressure and decreasing temperature

2. Using the Gibbs energy change,  $\Delta G^\circ = +63.3 \text{ kJ}$  for the following reaction,

- (a)  $3.2 \times 10^{-26}$
- (b)  $8.0 \times 10^{-12}$
- (c)  $2.9 \times 10^{-3}$
- (d)  $7.9 \times 10^{-7}$

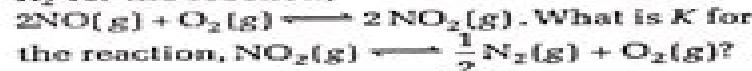
3.  $KMnO_4$  can be prepared from  $K_2MnO_4$  as per reaction.The reaction can go to completion by removing  $OH^-$  ions by adding [2013]

- (a) HCl
- (b) KOH
- (c)  $CO_2$
- (d)  $SO_2$

4. The value of  $\Delta H$  for the reaction,is less than zero. Formation of  $XY_4(g)$  will be favoured at [2011]

- (a) low pressure and low temperature
- (b) high temperature and low pressure
- (c) high pressure and low temperature
- (d) high temperature and high pressure

5. For the reaction,



- (a)  $1/(4K_1K_2)$
- (b)  $(1/K_1K_2)^{1/2}$
- (c)  $1/(K_1K_2)$
- (d)  $1/(2K_1K_2)$

6. In which of the following equilibrium  $K_c$  and  $K_p$  are not equal? [2010]

- (a)  $2NO(g) \rightleftharpoons N_2(g) + O_2(g)$
- (b)  $SO_2(g) + NO_2(g) \rightleftharpoons SO_3(g) + NO(g)$
- (c)  $H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$
- (d)  $2C(s) + O_2(g) \rightleftharpoons 2CO_2(g)$

7. The dissociation constants for acetic acid and HCN at  $25^\circ\text{C}$  are  $1.6 \times 10^{-5}$  and  $4.5 \times 10^{-10}$ , respectively. The equilibrium constant for the equilibrium,

- (a)  $3.0 \times 10^5$
- (b)  $3.0 \times 10^{-5}$
- (c)  $3.0 \times 10^4$
- (d)  $3.0 \times 10^6$

# Chemical Equilibrium Chemistry Study Guide Answers

**Steven S. Zumdahl, Paul B. Kelter**

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