

16 • Chemical Equilibrium

1. Consider the reaction system.
 $\text{CoO}(s) + \text{H}_2(g) \rightleftharpoons \text{Co}(s) + \text{H}_2\text{O}(g)$.

The equilibrium constant expression is

a) $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$ d) $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$
 b) $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$ c) $\frac{[\text{H}_2\text{O}]}{[\text{H}_2]}$
 (Handwritten note: $\text{CoO} \rightarrow \text{Co}$)
 (Handwritten note: $\text{H}_2 \rightarrow \text{H}_2\text{O}$)

2. Given the equilibrium,
 $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$, if this equilibrium is established by beginning with equal number of moles of SO_2 and O_2 in a 1.0 Liter bulb, then the following *must* be true at equilibrium:
 a) $[\text{SO}_2] = [\text{SO}_3]$ d) $[\text{SO}_2] < [\text{O}_2]$
 (Handwritten note: $2[\text{SO}_2] = 2[\text{SO}_3]$) e) $[\text{SO}_2] > [\text{O}_2]$
 c) $[\text{SO}_2] = [\text{O}_2]$

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole NO , 0.200 mol Cl_2 and 0.500 mol ClNO were placed in a 25.0 Liter container. The following equilibrium is established: $2\text{ClNO}(g) \rightleftharpoons 2\text{NO}(g) + \text{Cl}_2(g)$

3. At equilibrium, 0.600 mol of ClNO was present. The number of *moles* of Cl_2 present at equilibrium is
 a) 0.050 (d) 0.200
 b) 0.100 e) 0.250
 c) 0.150

PRACTICE TEST

4. The equilibrium constant, K_c , is:
 a) 4.45×10^{-4} d) 0.167
 b) 6.67×10^{-4} e) 1500
 (Handwritten note: $\text{c) } 0.111$)

5. At 985°C , the equilibrium constant for the reaction,
 $\text{H}_2(g) + \text{CO}_2(g) \rightleftharpoons \text{H}_2\text{O}(g) + \text{CO}(g)$

is 1.63. What is the equilibrium constant for the reverse reaction?
 a) 1.63 d) 0.613
 b) 0.815 (Handwritten note: $\text{c) } 1.00$)
 c) 2.66

6. What is the relationship between K_p and K_c for the reaction, $2\text{ICl}(g) \rightleftharpoons \text{I}_2(g) + \text{Cl}_2(g)$?
 a) $K_p = K_c(RT)^{-1}$ d) $K_p = K_c$
 b) $K_p = K_c(RT)$ e) $K_p = K_c(2RT)$
 c) $K_p = K_c(RT)^2$

7. For the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$, K_p at 25°C is 7.3, when all partial pressures are expressed in atmospheres. What is K_c for this reaction? [$R=0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$]
 (Handwritten note: $\text{a) } 4270$) d) 179
 b) 0.0119 e) 2.06
 c) 0.291

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About the Author Thomas A Evangelist teaches high school science for the New York City Board of Education

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