

# 16 • Chemical Equilibrium

1. Consider the reaction system,



The equilibrium constant expression is

- a)  $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$  d)  $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$   
b)  $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$  e)  $\frac{[\text{H}_2\text{O}]}{[\text{H}_2]}$   
c)  $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{H}_2]}$

2. Given the equilibrium,



if this equilibrium is established by beginning with equal number of moles of  $\text{SO}_2$  and  $\text{O}_2$  in a 1.0 Liter bulb, then

the following *must* be true at equilibrium:

- a)  $[\text{SO}_2] = [\text{SO}_3]$  d)  $[\text{SO}_2] < [\text{O}_2]$   
b)  $2[\text{SO}_2] = 2[\text{SO}_3]$  e)  $[\text{SO}_2] > [\text{O}_2]$   
c)  $[\text{SO}_2] = [\text{O}_2]$

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole  $\text{NO}$ , 0.200 mol  $\text{Cl}_2$  and 0.500 mol  $\text{ClNO}$  were placed in a 25.0 Liter container. The following equilibrium is established:  $2\text{ClNO(g)} \rightleftharpoons 2\text{NO(g)} + \text{Cl}_2\text{(g)}$

3. At equilibrium, 0.600 mol of  $\text{ClNO}$  was present. The number of *moles* of  $\text{Cl}_2$  present at equilibrium is

- a) 0.050 d) 0.200  
b) 0.100 e) 0.250  
c) 0.150

## PRACTICE TEST

4. The equilibrium constant,  $K_c$ , is:

- a)  $4.45 \times 10^{-4}$  d) 0.167  
b)  $6.67 \times 10^{-4}$  e) 1500  
c) 0.111

5. At  $985^\circ\text{C}$ , the equilibrium constant for the reaction,



is 1.63. What is the equilibrium constant for the reverse reaction?

- a) 1.63 d) 0.613  
b) 0.815 e) 1.00  
c) 2.66

6. What is the relationship between  $K_p$  and  $K_c$  for the reaction,  $2\text{HCl(g)} \rightleftharpoons \text{H}_2\text{(g)} + \text{Cl}_2\text{(g)}$ ?

- a)  $K_p = K_c(RT)^{-1}$  d)  $K_p = K_c$   
b)  $K_p = K_c(RT)$  e)  $K_p = K_c(2RT)$   
c)  $K_p = K_c(RT)^2$

7. For the reaction  $2\text{NO}_2\text{(g)} \rightleftharpoons \text{N}_2\text{O}_4\text{(g)}$ ,  $K_p$  at  $25^\circ\text{C}$  is 7.3, when all partial pressures are expressed in atmospheres. What is  $K_c$  for this reaction? [ $R=0.0821 \text{ L}\cdot\text{atm}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ ]

- a) 4270 d) 179  
b) 0.0119 e) 2.06  
c) 0.291

# Chemical Equilibrium Practice Test

**Kaplan**

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