

Chapter 18

Study Guide - Answers

1. Write the expression for the equilibrium constant for each of the following reactions.



$$K_{eq} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$



$$K_{eq} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$



$$K_{eq} = \frac{[\text{H}^+][\text{CN}^-]}{[\text{HCN}]}$$



$$K_{eq} = \frac{[\text{H}^+][\text{HSO}_4^-]}{[\text{H}_2\text{SO}_4]}$$



$$K_{eq} = \frac{[\text{Pb}^{2+}][\text{F}^-]^2}{}$$

2. From the data provided below, calculate the value of the equilibrium constant for the reaction.



$$[\text{H}_2] = [\text{Cl}_2] = 1.0 \times 10^{-2}; [\text{HCl}] = 1.0 \times 10^{-4}$$

$$K_{eq} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$

$$K_{eq} = \frac{[1.0 \times 10^{-4}]^2}{[1.0 \times 10^{-2}][1.0 \times 10^{-2}]} = 0.0001$$



$$[\text{N}_2] = 4.4 \times 10^{-2}; [\text{H}_2] = 1.2 \times 10^{-1}; [\text{NH}_3] = 3.4 \times 10^{-3}$$

$$K_{eq} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$

$$K_{eq} = \frac{[3.4 \times 10^{-3}]^2}{[4.4 \times 10^{-2}][1.2 \times 10^{-1}]^3} = 0.152$$

3. For the reaction $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$, $K_{eq} = 0.2$. At a particular time, the following concentrations are measured: $[\text{N}_2\text{O}_4] = 2.0 \text{ M}$, $[\text{NO}_2] = 0.2 \text{ M}$. Is this reaction at equilibrium? If not, in which direction will the reaction proceed?



$$[\text{N}_2\text{O}_4] = 2.0 \text{ M}$$

$$[\text{NO}_2] = 0.2 \text{ M}$$

$$Q = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$$

$$Q = \frac{[0.2]^2}{[2.0]} = 0.02$$

$$K_{eq} = 0.2$$

$Q < K_{eq}$, so the reaction will proceed to the right (forward).

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Daniel L. Reger



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Table of Contents Chemical Equilibrium Study Guide Answers

1. Understanding the eBook Chemical Equilibrium Study Guide Answers
 - The Rise of Digital Reading Chemical Equilibrium Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equilibrium Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Equilibrium Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Equilibrium Study Guide Answers
 - Personalized Recommendations
 - Chemical Equilibrium Study Guide Answers User Reviews and Ratings
 - Chemical Equilibrium Study Guide Answers and Bestseller Lists
5. Accessing Chemical Equilibrium Study Guide Answers Free and Paid eBooks
 - Chemical Equilibrium Study Guide Answers Public Domain eBooks
 - Chemical Equilibrium Study Guide Answers eBook Subscription Services
 - Chemical Equilibrium Study Guide Answers Budget-Friendly Options
6. Navigating Chemical Equilibrium Study Guide Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Chemical Equilibrium Study Guide Answers Compatibility with Devices
 - Chemical Equilibrium Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chemical Equilibrium Study Guide Answers
 - Highlighting and Note-Taking Chemical Equilibrium Study Guide Answers
 - Interactive Elements Chemical Equilibrium Study Guide Answers
8. Staying Engaged with Chemical Equilibrium Study Guide Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Equilibrium Study Guide Answers
9. Balancing eBooks and Physical Books Chemical Equilibrium Study Guide Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Equilibrium Study Guide Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemical Equilibrium Study Guide Answers
- Setting Reading Goals Chemical Equilibrium Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Equilibrium Study Guide Answers
- Fact-Checking eBook Content of Chemical Equilibrium Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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- Integration of Multimedia Elements
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